

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

\_\_\_\_\_ **lattice energy** \_\_\_\_\_

a. ionic bond

\_\_\_\_\_ **bond energy** \_\_\_\_\_

b. covalent bond

\_\_\_\_\_ **enthalpy of vaporization** \_\_\_\_\_

c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

\_\_\_\_\_ **ionic** \_\_\_\_\_

a. H—F

\_\_\_\_\_ **nonpolar covalent** \_\_\_\_\_

d. H—H

\_\_\_\_\_ **ionic** \_\_\_\_\_

b. Na—Cl

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

e. H—C

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

c. H—O

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen  
atom and the unshared electron pair of a strongly electronegative atom such as O,  
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,  
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally  
intermolecular, while ionic and covalent bonds occur between ions or atoms  
respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.  
Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making  
water a liquid at room temperature. Hydrogen bonding exists between water  
molecules, but not between hydrogen sulfide molecules.

# Review Chemical Bonding Section 5 Answers

**L Darling-Hammond**



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