

## Stoichiometry Review

1. The combustion of propene is represented:  

$$\text{C}_3\text{H}_6\text{g} + 5\text{O}_2\text{g} \rightarrow 3\text{CO}_2\text{g} + 3\text{H}_2\text{Og}$$
 How many grams of oxygen are required to burn completely 10.0 g of propene?
2. Small quantities of sodium can be prepared in the laboratory by the electrolysis:  

$$\text{NaCl(aq)} + \text{H}_2\text{O(l)} \rightarrow \text{NaOH(aq)} + \text{H}_2\text{g} + \text{Cl}_2\text{g}$$
 What volume of chlorine at STP would be prepared from 10.0 g of sodium chloride?
3. The heat content of ethane is 850 kJ/mol.  $\text{C}_2\text{H}_6\text{g}$ . Upon heating this gas, 1000 kJ/mol is obtained. How many grams of  $\text{C}_2\text{H}_6\text{g}$  are required to produce 100 g of ethene, according to the following reaction?  

$$2\text{C}_2\text{H}_6\text{g} \rightarrow \text{C}_2\text{H}_4\text{g} + \text{H}_2\text{g}$$
4. Ethanol is oxidized in a charge transfer by the permanganate ion. The overall reaction in the acidic is:  

$$5\text{C}_2\text{H}_5\text{OHg} + 4\text{H}^+\text{g} + 2\text{MnO}_4^-\text{g} \rightarrow 5\text{C}_2\text{H}_4\text{Og} + 2\text{Mn}^{2+}\text{g}$$
 Calculate the volume of oxygen at STP required to convert 100.0 g of glucose to carbon dioxide and water.
5. Acetylene is used in welding and is prepared by the reaction of calcium carbide with water:  

$$\text{CaC}_2 + 2\text{H}_2\text{O} \rightarrow \text{C}_2\text{H}_2 + \text{Ca(OH)}_2$$
 How many kilograms of calcium carbide will be produced from 100.0 kg of acetylene chloride and HCl by the following balanced chemical equation?
6. Bromine can be prepared by adding chlorine to a solution of sodium bromide. How many grams of bromine are formed if 100.0 g of chlorine reacts with an excess of sodium bromide in the reaction below. What percentage of chlorine is converted to bromine?  

$$2\text{NaBr(aq)} + \text{Cl}_2\text{(g)} \rightarrow 2\text{NaCl(aq)} + \text{Br}_2\text{(l)}$$
7. Aqueous solutions of  $\text{HNO}_3$  and  $\text{KOH}$  are mixed and reacted to form  $\text{KNO}_3$  and  $\text{H}_2\text{O}$  according to the following reaction:  

$$\text{KOH} + \text{HNO}_3 \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$$
8.  $\text{H}_2\text{SO}_4$  is a sample of sulphuric acid is neutralized with an excess of sodium hydroxide. Calculate the mass of phosphorus pentoxide produced.
9. How many grams of oxygen are prepared by the decomposition of 100.0 g of hydrogen peroxide?
10. What volume of oxygen gas at STP can be prepared by the decomposition of 100.0 g of potassium chlorate?
11. How many grams of zinc are required for the replacement of  $\text{CuSO}_4$  and hydrogen from a solution of  $\text{ZnSO}_4$ .
12. What mass of magnesium chloride is required to react with  $\text{HCl}$  and silver nitrate?

# Review Chemistry 1 Honors Unit 10 Stoichiometry

**Stanford University**



## **Review Chemistry 1 Honors Unit 10 Stoichiometry:**

*Stanford University Bulletin* Stanford University, 1999      General Catalog Iowa State University, 1987      **Formulation and Stoichiometry** Emil J. Margolis, 2012-12-06

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