Exam 2 Review (Chapters 12 and 13)

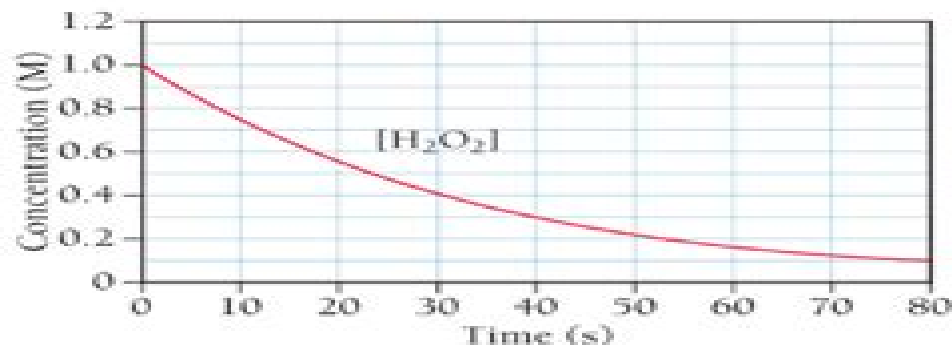1. Consider the balanced equation below

   $$CH_4 + 4Cl_2 \longrightarrow CCl_4 + 4HCl$$

   Write the rate of the following reaction in terms of the appearance of products and disappearance of reactants.

2.

The graph above shows the concentration of H_2O_2 as a function of time for the reaction $2H_2O_2 (aq) \longrightarrow 2H_2O(l) + O_2 (g)$. Using the graph complete the Following.

   a) Find the average rate of the reaction between 10 and 20 seconds

Review Module Chapters 12 Answer

K Morrison

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