

**WEST BENGAL COUNCIL OF HIGHER SECONDARY EDUCATION**  
**SYLLABUS FOR CLASSES XI AND XII**  
**SUBJECT : CHEMISTRY (CHEM)**

**CLASS - XI**

**SEMESTER – I**

**SUBJECT : CHEMISTRY (CHEM)**

**FULL MARKS : 35**

**CONTACT HOURS : 70 Hours**

**COURSE CODE : THEORY**

**Sub-topics**

UNIT No.	TOPICS	CONTACT HOURS	MARKS
Unit - 1	<b>Some Basic Concepts of Chemistry:</b> Laws of chemical combination. Concept of elements, atoms and molecules. Atomic and molecular masses. Mole concept and molar mass, percentage composition, empirical and molecular formula, chemical reactions, stoichiometry and calculations based on stoichiometry. Different concentration terms of solutions and related calculations.	07	03
Unit - 2	<b>Structure of Atom:</b> Bohr's model and its limitations, concept of shell and sub-shells, the dual nature of matter and light, de Broglie's relationship. Heisenberg uncertainty principle, Schrödinger wave equation (elementary idea only). Concept of orbitals, quantum numbers, shapes of s, p and d orbitals, rules for filling electrons in orbitals: Aufbau principle, Pauli exclusion principle and Hund's rule, exchange energy, electronic configuration of atom, stability of half-filled, completely filled orbitals.	12	06
Unit - 3	<b>Classification of Elements and Periodicity in Properties:</b> Modern periodic law and the present form of the periodic table, periodic trends in properties of elements – atomic radii, ionic radii, van der Waals' radii, ionization enthalpy, electron gain enthalpy, electronegativity, valency. Nomenclature of elements with atomic number greater than 100.	07	04
Unit - 4	<b>Chemical Bonding and Molecular Structure:</b> Valence electrons, ionic bond, bond parameters, covalent bond, Lewis structure, polar character of covalent bond, covalent character of ionic bond, valence bond theory, resonance, geometry of covalent molecules, VSEPR theory, concept of hybridisation, involving s, p and d orbitals and shapes of some simple molecules, intermolecular interactions, Hydrogen bonding. Molecular orbital theory of homonuclear diatomic molecules ( $H_2$ , $He_2$ , $O_2$ , $N_2$ , $F_2$ – qualitative idea only)	13	06

# West Bengal Board Class Xii Science Syllabus

**Lei Shi**



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 function has the limit say  $c$  as  $(x, y) \rightarrow (a, b)$  provided that  $\epsilon > 0$  there exists  $\Delta > 0$  such that  $|f(x, y) - c| < \epsilon$  whenever  $0 < \Delta < \Delta_0$

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 to 2 mbox implies y 2 9 end align begin align mathop  $\lim_{y \rightarrow 2} y$  to 2 g left y right mathop  $\lim_{y \rightarrow 2} y$  to 2 3

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 as we get closer and closer to  $x = 3$

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 the limit of  $f$  as  $x$  approaches  $a$  is  $l$  3 how do you read  $\lim_{x \rightarrow a} f(x)$  solution the limit of  $f$  as  $x$  approaches  $a$  from the left 4 how

do you read  $\lim_{x \rightarrow a} f(x)$  solution the limit of  $f$  as  $x$  approaches

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$$\lim_{x \rightarrow a} c = c \quad \lim_{x \rightarrow a} x = a \quad \lim_{x \rightarrow a} c x = c \lim_{x \rightarrow a} x = c a$$

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