

A 23.8 L balloon is filled with 88g CO_2
at 15°C , what is the pressure in

kPa?

$$PV = nRT$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$PV = nRT$$

$$P(23.8\text{L}) = (2)$$

$$V = \underline{23.8\text{L}}$$

$$T = 15^\circ\text{C} + 273 = \underline{288\text{K}}$$

$$P = ? \text{ kPa}$$

$$88\text{g CO}_2 \times \frac{1\text{mol CO}_2}{(12 + 16 \times 2)} = \underline{2\text{mol}}$$

44g

Weston Walch Gas Law Problems

JS Bruner



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