

Week 10: CHEM 2700 -- Unit 4 - Reaction Rates and Equilibrium

Unit 4: Reaction Rates and Equilibrium - Chemistry is everywhere.

PRACTICE 2 - Equilibrium Expressions

1. Write an equilibrium expression for each of the following chemical reactions.



$$K_c = \frac{[\text{N}_2\text{O}_4]}{[\text{NO}_2]^2}$$



$$K_c = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$



$$K_c = \frac{[\text{NO}_2]^2 [\text{Cl}_2\text{O}] [\text{Cl}_2]}{[\text{NO}_2\text{Cl}]^2}$$



$$K_c = \frac{[\text{NO}_2]^2 [\text{Cl}_2]}{[\text{NO}_2\text{Cl}]^2}$$



$$K_c = \frac{[\text{N}_2\text{O}_4] [\text{Cl}_2]^2}{[\text{NO}_2\text{Cl}_2]^2}$$

2. The following equilibria are established between substances. Write an equilibrium expression for each reaction.



$$K_c = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}] [\text{SCN}^-]}$$



$$K_c = \frac{[\text{AgClO}_3]}{[\text{Ag}^+] [\text{ClO}_3^-]}$$



$$K_c = \frac{[\text{AgSCN}]}{[\text{Ag}^+] [\text{SCN}^-]}$$

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Hussin A.Rothana



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Scientific and Technical Aerospace Reports ,1994 *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science* ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

Bulletin of the Atomic Scientists ,1972-10 The Bulletin of the Atomic Scientists is the premier public resource on scientific and technological developments that impact global security Founded by Manhattan Project Scientists the Bulletin s iconic

Doomsday Clock stimulates solutions for a safer world **Physical Chemistry of Foods** Pieter Walstra,2002-10-08
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