

DETERMINATION

DETERMINATION

1. The first step in the determination of the concentration of a substance in a sample is to weigh a known amount of the sample. This is done by using a balance and a weighing boat. The weight of the sample is then measured and recorded.

2. The next step is to dissolve the sample in a known volume of solvent. This is done by adding the sample to a volumetric flask and filling it to the mark with the solvent. The volume of the solution is then measured and recorded.

3. The third step is to measure the absorbance of the solution. This is done by using a spectrophotometer. The absorbance is then measured at a specific wavelength and recorded.

4. The fourth step is to calculate the concentration of the substance in the sample. This is done by using the Beer-Lambert law, which relates absorbance to concentration and path length.

5. The final step is to report the concentration of the substance in the sample. This is done by expressing the concentration in units of mass per volume (e.g., mg/L).

6. The concentration of the substance in the sample is then compared to the concentration of the substance in the standard solution. This is done by using a calibration curve, which relates absorbance to concentration.

7. The concentration of the substance in the sample is then reported as a percentage of the concentration of the substance in the standard solution.

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29. The concentration of the substance in the sample is then reported as a percentage of the concentration of the substance in the standard solution.

Concentration (mg/L)	Absorbance
0	0.000
10	0.150
20	0.300
30	0.450
40	0.600
50	0.750
60	0.900
70	1.050
80	1.200
90	1.350
100	1.500

Concentration (mg/L)	Absorbance
0	0.000
10	0.150
20	0.300
30	0.450
40	0.600
50	0.750
60	0.900
70	1.050
80	1.200
90	1.350
100	1.500

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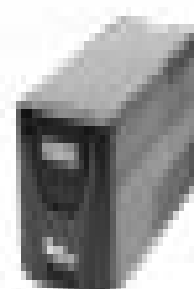
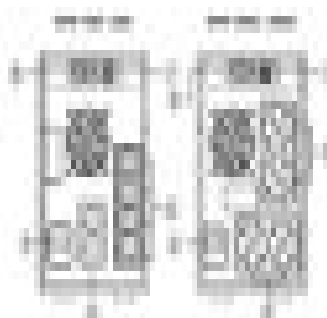
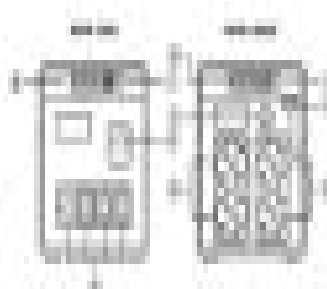
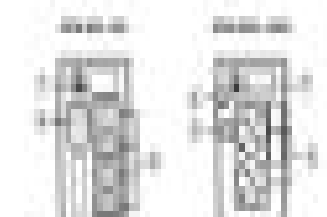
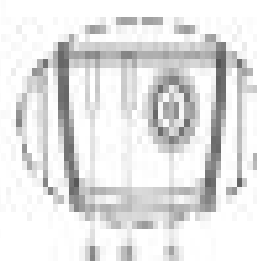


Figure 1



Figure 2



Riello Ups User Manual

J Elliott



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