- 6. Write a pair of balanced half-reaction ocusions—one showing a gair of electrons and one showing a loss of electrons - for each of the following receivers:

  - (a)  $Z \gamma_{00} + Cal_{eq} \rightarrow E \eta_{eq}^{-} + Cal_{eq}$ (b)  $M Q_{00} + 2 H_{eq}^{-} \rightarrow M \eta_{eq}^{-} + F_{eq}$

600 Chapter 9

Sections 90.5

- 9. For each of the following, write and label the oxidation and reduction half-reaction equations, ignore speciation ions.
  - 66 Night + CHENDy Joseph -> Corp. + Night Corp.
  - (b) Play + Ca(NCa)read -> Cupe + Pb(NO<sub>3</sub>)read
  - (c)  $Ga_{(a)} + 2HNO_{X(a)} \rightarrow H_{X(a)} + G_{X}(NG_{A})_{X(a)}$
  - (d)  $2M_{\rm bl} = {\rm Fe}_{\rm i}O_{\rm bk} \rightarrow 2{\rm Fe}_{\rm Cl} = M_{\rm i}O_{\rm bkl}$  (Figure 8)
- 18. We now only orded at one type of single deplacement reaction—a metal displacing another metal from an lonic compound. A nonmetal can also displace enother renmetal from an ionic compound. For example,

Using your knowledge of atoms and ions and the ideas presented in this shapter, write a pair of halanced half-warting or surjeys for this ownition—one showing a pain of electrons and one showing a loss of electrons



Figure 8. The restuction of ston(IR) colds for atuminum is called the "thermine" reaction. Documes this reaction is

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