

# CHEMISTRY Terminology

The ABCs of chemistry

## A

**A:** Symbol for atomic mass number  
**absolute zero:** Lowest temperature; 0 Kelvin  
**Ae:** Symbol for the element actinium,  $Z = 89$   
**accuracy:** The closeness of measurement to true value  
**acetate anion:**  $C_2H_3O_2^-$ , charge -1  
**acetic acid:**  $C_2H_4O_2$   
**acetone:** Dimethyl ketone; organic solvent  
**acid-base indicator:** A chemical that changes color if the pH changes  
**acid-base titration:** A method for determining acid or base concentrations  
**acid definition:** See Arrhenius acid, Brønsted-Lowry acid, Lewis acid  
**acid ionization constant ( $K_a$ ):** Equilibrium constant for acid dissociation  
**acidic solution:** pH below 7  
**actinide:** Element with  $Z = 90$  or above; radioactive  
**activation energy ( $E_a$ ):** A process's energy barrier



**adhesion:** Attraction of unlike molecules or materials  
**Ag:** Symbol for the element silver,  $Z = 47$   
**Al:** Symbol for the element aluminum,  $Z = 13$   
**alcohol:** Organic compound with -OH group; ROH  
**aldehyde:** Carbonyl group bonded to hydrogen and an organic group  
**alkali metal family:** Lithium, sodium, potassium, rubidium, cesium, and francium; column #1  
**alkaline earth metal family:** Beryllium, magnesium, calcium, strontium, barium, and radium; column #2  
**alkane:** Hydrocarbon; all C-C single bonds  
**alkene:** Hydrocarbon; 1 or more C=C double bonds  
**alkyne:** Hydrocarbon; 1 or more C≡C triple bonds  
**allotropes:** Two or more forms of an element  
**alloy:** Solution of 2 or more metals  
**alpha ( $\alpha$ ):** Greek letter that denotes radioactive particle and various scientific variables  
**alpha particle:** Helium nucleus; charge +2  
**Au:** Symbol for the element auriferous,  $Z = 79$   
**amalgam:** Alloy of mercury with other metal  
**amine:** Organic base; RRRN; modified ammonia  
**amino acid:** Compound with organic acid and organic base property; forms proteins and peptides  
**ammonia:**  $NH_3$ , base  
**ammonium cation:**  $NH_4^+$ , charge +1  
**ammonium hydroxide:**  $NH_4OH$   
**amphoteric oxide:** Exhibits both acid and basic properties  
**amu:** Atomic mass unit;  $1/12$  mass of C-12  
**anion:** Ion with a negative charge  
**anode:** Electrode that supports oxidation  
**antibonding MO:** MO is less stable than separate AOs  
**AO:** Abbreviation for atomic orbital

**aq:** Aqueous; "dissolved in water"  
**Ar:** Symbol for the element argon,  $Z = 18$   
**aromatic:** Organic compound with a benzene ring  
**Arrhenius acid:** Produces hydronium ion in water solution  
**Arrhenius base:** Yields hydroxide ion in water solution  
**Arrhenius equation for rate constant ( $k$ ):**  

$$k = Ae^{-E_a/RT}$$
 $E_a$  = activation energy  
**arsenide anion:**  $As^{3-}$ , charge -3  
**As:** Symbol for the element arsenic,  $Z = 33$   
**At:** Symbol for the element astatine,  $Z = 85$   
**atm:** Symbol for pressure in "atmospheres"  
**atom:** Fundamental unit of all matter  
**atomic mass number ( $A$ ):** Total number of protons and neutrons  
**atomic number ( $Z$ ):** Number of protons in the nucleus  
**atomic orbital:** Wave motion of electrons in atoms  
**atomic radius:** Empirical measure of atom size  
**atomic weight:** Weighted average of natural isotopes of an element  
**Au:** Symbol for the element gold,  $Z = 79$   
**Aufbau principle:** Guides the filling of electronic subshells of the elements  
**Avogadro's law:** Volume is proportional to gas moles at fixed pressure and temperature  
**Avogadro's number:**  $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

**Bk:** Symbol for the element berkelium,  $Z = 97$   
**bleach:** Chemical that is a strong oxidizing agent  
**boiling:** Liquid  $\rightarrow$  gas at the boiling point  
**boiling point elevation:** A solution has a higher boiling point than a pure solvent  
**boiling point ( $T_b$ ):** Liquid-gas equilibrium,  $P = 1 \text{ atm}$   
**boiling point of water:** 100°C  
**bond energy:** Energy held by a chemical bond  
**bond length:** Distance between 2 atoms in a bond  
**bond order for Lewis model:** # of bonds divided by # of bonded atoms  
**bond order for MO treatment:** # of filled bonding MOs minus # of filled antibonding MOs  
**bonding electrons:** Form chemical bonds in a compound  
**bonding MO:** MO is more stable than AOs  
**Boyle's law:**  $PV$  constant for gas, fixed  $T$  and  $n$   
**Br:** Symbol for the element bromine,  $Z = 35$   
**bromide anion:**  $Br^-$ , charge -1  
**Brønsted-Lowry acid:** Proton donor  
**Brønsted-Lowry base:** Proton acceptor  
**buffer:** Composed of weak acid and weak base; serves to keep pH constant



## B

**B:** Symbol for the element boron,  $Z = 5$   
**Ba:** Symbol for the element barium,  $Z = 56$   
**balanced equation:** Each side of the equation has the same number of atoms of each element and charge  
**balanced redox:** Electron loss = electron gain  
**barium cation:**  $Ba^{2+}$ , charge +2  
**base definition:** See Arrhenius base, Brønsted-Lowry base, Lewis base  
**base ionization constant ( $K_b$ ):** Equilibrium constant for base dissociation  
**basic solution:** pH over 7  
**battery:** Source of electrical power; galvanic cell  
**Be:** Symbol for the element beryllium,  $Z = 4$   
**benzene:**  $C_6H_6$ ; organic solvent; aromatic ring of 6 carbon atoms  
**beryllium cation:**  $Be^{2+}$ , charge +2  
**beta ( $\beta$ ):** Greek letter that denotes radioactive particle and a number of scientific variables  
**beta particle:** Energetic electron from the nucleus; charge -1  
**Bh:** Symbol for the element bohrium,  $Z = 107$   
**Bi:** Symbol for the element bismuth,  $Z = 83$   
**bicarbonate anion:**  $HCO_3^-$ , charge -1  
**bimolecular oxygen:**  $O_2$   
**binary compound:** Formed from 2 elements  
**bisulfate anion:**  $HSO_4^-$ , charge -1



**b:** Symbol for  $\text{centi-}$ , the SI prefix for  $1/100$  or  $10^{-2}$   
**c:** Symbol for the speed of light  
**°C:** Celsius temperature scale  
**C:** Symbol for the element carbon,  $Z = 6$   
 **$C_2H_4O_2$ :** Acetic acid  
**Ca:** Symbol for the element calcium,  $Z = 20$   
**calcium cation:**  $Ca^{2+}$ , charge +2  
**calcium fluoride:**  $CaF_2$   
**calcium hydroxide:**  $Ca(OH)_2$   
**calorie:** English unit of heat energy  
**calorimetry:** Study of the release or absorption of heat  
**carbide anion:**  $C^{4-}$ , charge -4  
**carbohydrate:** Organic compound; carbon bonded to several -OH groups (sugar and starch)  
**carbon dioxide:**  $CO_2$   
**carbon monoxide:**  $CO$   
**carbonate anion:**  $CO_3^{2-}$ , charge -2  
**carbonic acid:**  $H_2CO_3$   
**carbonyl group:**  $CO$ , ligand complex  
**carboxyl group:**  $-CO$  bonded to 2 groups  
**carboxylic acid:** Organic acid;  $R-COOH$   
**catalyst:** Accelerates a reaction but is not consumed in the reaction; lowers the activation energy  
**cathode:** Electrode that supports reduction  
**cation:** Ion with a positive charge  
**Cd:** Symbol for the element cadmium,  $Z = 48$



# Sol Chemistry Study Guide

**Scott C. Dulebohn**



## **Sol Chemistry Study Guide:**

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Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Cracking the Virginia SOL* Tina Walton, Princeton Review (Firm), 2000 The Princeton Review realizes that acing the English Reading Literature and Research exam is very different from getting straight As in school The Princeton Review doesn't try to teach students everything there is to know about English only the techniques they'll need to score higher on the exam There's a big difference In *Cracking the Virginia SOL EOC English Reading Literature and Research* The Princeton Review will teach test takers how to think like the test makers and Eliminate incorrect answer choices by using Process of Elimination and other techniques Familiarize students with every type of reading material that might be on the test Use the two pass system to get a grip on the test structure and raise scores Hone individual skills with practice questions for each type of test question This book includes 2 full length simulated exams All of TPR's sample test questions are just like the ones test takers will see on the actual End of Course English Reading Literature and Research exam and TPR fully explains every solution Contents Include The Mystery Test Structure and Strategies II Terms Literary Movements and Writers Literature Review III Literary Elements and Printed and Resource Materials The Elements of Literature A Variety of Printed Material A Variety of Resource Material IV The Princeton Review Practice Tests **A Self-study Guide to the Principles of Organic Chemistry** Jiben Roy, 2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration

bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix     Self Study Guide CLAT 2022 Arihant Experts,2021-10-14 1 The Self Study Guide for the preparation of CLAT 2021 2 The book is divided into 6 Sections 3 Good number of MCQs have been provided for practice 4 Chapterwise Previous Years Papers and 3 Crack sets are given for thorough practice 5 Authentic Solutions are given for the complete assistance The Self Study Guide CLAT 2021 is a comprehensive textbook designed to give complete assistance for the preparation This book divides the entire syllabus into sections with Chapterwise theories along with sufficient number of MCQs are given for quick revision of topics making it a complete success package Each chapter carries previous years question from 2017 to 2011 for practice and 3 Practice sets are provided at end of the book to analyze the level of preparation With authentic solutions provided help students to strengthen the concepts Students who desire to score well in CLAT this self study guide is a perfect choice for them Toc CLAT Solved Paper 2020 2017 Legal Aptitude Indian Constitution English Language Mathematics Logical Reasoning General Knowledge Practice Sets 1 3     A Study Guide for Jose Raul Bernardo's "Silent Wing" Gale, Cengage Learning,2016 A Study Guide for Jose Raul Bernardo s Silent Wing excerpted from Gale s acclaimed Literature of Developing Nations for Students This concise study guide includes plot summary character analysis author biography study questions historical context suggestions for further reading and much more For any literature project trust Literature of Developing Nations For Students for all of your research needs     *Study Guide to Accompany Calculus for the Management, Life, and Social Sciences* Clyde Metz,1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences     Self Study Guide CLAT 2020 Arihant Experts,2019-08-12 CLAT Common Law Admission Test is an All India entrance examination and conducted by the National Law Schools Universities for the admission in the Under graduate and Post graduate programmes This examination is given right after 12 th grade for Undergraduate programmes in Law and after graduation in Law for Masters of Law This Self Study Guide CLAT 2020 book comprehensively covers the whole syllabus regarding to test and helps students to get through exam without any difficulty This book is divided into 6 parts namely Legal Aptitude Indian Constitution English Language General Knowledge Logical Reasoning and Mathematics This book provides Chapter wise theories with adequate amount of MCQs for recalling the concepts of each chapter Previous Years question papers have been provided in the chapter practice 3 sets of practice papers are provided at the end of the book which describes the exact pattern of the question to the candidates so that they can answer easily detailed explanation of solved papers 2019 2018 also have been provided for practice All the practice sets are also available for free on the WEB as well as on MOBILE so that

candidates can prepare for their examination at anytime from anywhere This book highly useful for those candidates who willing to clear the exams with the good grades and planning to study from the great institutes TABLE OF CONTENT CLAT Solved Paper 2019 CLAT Solved Paper 2018 CLAT Solved Paper 2017 Legal Aptitude Indian Constitution English Language Mathematics Logical Reasoning and General Knowledge Practice Sets 1 3      *Organic Chemistry, Study Guide and Solutions Manual* T. W. Graham Solomons, Craig B. Fryhle, 2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry      *Self Study Guide B. Pharma Entrance Exam 2021* Arihant Experts, 2020-11-09 1 B Pharma Entrance Examination 2021 is a one point solution for the entrance exam 2 The book is divided into 4 sections 3 Previous Years Solved papers are given for the practice 4 Precise and detailed text with illustrations eases in learning the concepts 5 This book uses the easy language for better understanding Bachelor of Pharmacy B Pharma is a 4 years undergraduate program in which students study the methods and process of preparing medicines To get into the proper college or institution one needs to clear the entrance exam that tests the suitability and apparent knowledge required for the course The Self Study Guide of B Pharma Entrance Examination 2021 is an on point solution for various B Pharma Entrances conceived and designed as according to latest exam pattern Precise and detailed text with illustrations makes it suitable for all categories of students Strict approach towards the prescribed syllabus enables students to get focused preparation Also Last 9 Years Solved Papers are provided following the actual trends of the exams and helping students to get prepared accordingly A Must have book for those who really aspire to be a pharmacist TOC Solved Papers 2020 2012 Physics Chemistry Botany Zoology Appendix      Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application      **Introduction to Materials Chemistry** Harry R. Allcock, 2019-09-10 This textbook introduces the reader to the elementary chemistry on which materials science depends by discussing the different classes of materials

and their applications It shows the reader how different types of materials are produced why they possess specific properties and how they are used in technology Each chapter contains study questions to enable discussions and consolidation of the acquired knowledge The new edition of this textbook is completely revised and updated to reflect the significant expansion of the field of materials chemistry over the last years covering now also topics such as graphene nanotubes light emitting diodes extreme photolithography biomedical materials and metal organic frameworks From the reviews of the first edition This book is not only informative and comprehensive for a novice reader but also a valuable resource for a scientist and or an industrialist for new and novel challenges Materials and Manufacturing Process June 2009 Allcock provides a clear path by first describing basic chemical principles then distinguishing between the various major materials groups and finally enriching the student by offering a variety of special examples CHOICE April 2009 Proceeding logically from the basics to materials in advanced technology it covers the fundamentals of materials chemistry including principles of materials synthesis and materials characterization methods Internationale Fachzeitschrift Metall January 2009

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complete picture of the exam pattern and level both This book is highly useful and a complete guide for the aspirants those who are willing to make future in Hotel Management

**TABLE OF CONTENTS** Solved Papers 2020 2018 English Language and Comprehensive Reasoning and Logical Deduction Numerical Ability General Awareness Service Aptitude Group Discussion and Personal Interview Crack Sets 1 3 Answers to Section Tests and Crack Sets 1 3      **MEGA Study Guide for NTSE 2021 (SAT & MAT) Class 10 Stage 1 & 2 - 12th Edition** Disha Experts,2020-05-13 This new 12th edition of MEGA Study Guide for NTSE 2021 Class 10 is empowered with the inclusion of 2020 Stage I questions of the different states The book is based on the syllabus of Class 8 9 10 as prescribed by NCERT There are 27 chapters in the Mental Ability Section MAT The Scholastic Aptitude section SAT has been divided into 8 parts Physics Chemistry Biology Mathematics History Geography Civics and Economics The book provides past questions of last 10 years of NTSE Stage 1 2012 2020 Stage 2 2012 2019 JSTSE papers divided chapter wise The book provides sufficient pointwise theory solved examples followed by Fully Solved exercises in 2 levels State UT level National level Maps Diagrams and Tables to stimulate the thinking ability of the student The book covers new variety of questions Passage Based Assertion Reason Matching Definition based Statement based Feature Based Diagram Based and Integer Answer Questions      *MEGA Study Guide for NTSE (SAT, MAT & LCT) Class 10 Stage 1 & 2 - 11th Edition* Disha Experts,2019-03-12 This new 11th edition of MEGA Study Guide for NTSE Class 10 is empowered with the inclusion of 2018 Stage I questions of the different states The book is based on the syllabus of Class 8 9 10 as prescribed by NCERT The book also comprises of Past questions of NTSE Stage 1 2 from the years 2012 2018 There are now 28 chapters in the Mental Ability Section MAT The Scholastic Aptitude section SAT has been divided into 9 parts Physics Chemistry Biology Mathematics English History Geography Civics and Economics The book provides past questions of last 10 years of NTSE Stage 1 2 JSTSE papers divided chapter wise The book provides sufficient pointwise theory solved examples followed by Fully Solved exercises in 2 levels State UT level National level Maps Diagrams and Tables to stimulate the thinking ability of the student The book covers new variety of questions Passage Based Assertion Reason Matching Definition based Statement based Feature Based Diagram Based and Integer Answer Questions      **Chemistry & Cdr & Study Guide & Sol Manual & Stud Companion & Media Activities B** University Loretta Jones,University Lecturer in Physical Chemistry Fellow and Tutor P W Atkins,2000-05-30      *Oswaal CBSE Question Bank Class 12 English Core, Physics, Chemistry & Mathematics (Set of 4 Books) Chapterwise and Topicwise Solved Papers For Board Exams 2025* Oswaal Editorial Board,2024-02-15 Description of the product 100% Updated Syllabus Fully Solved Board Papers we have got you covered with the latest and 100% updated curriculum Crisp Revision with Topic wise Revision Notes Smart Mind Maps Extensive Practice with 3000 Questions Board Marking Scheme Answers to give you 3000 chances to become a champ Concept Clarity with 1000 Concepts 50 Concept Videos for you to learn the cool way with videos and mind blowing concepts NEP 2020 Compliance with Competency Based Questions for you to be on the cutting edge of the coolest educational trends

**Subject Guide to Books in Print** ,1997

Student Guide to the SAT David M. Tarlow,1987



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## Table of Contents Sol Chemistry Study Guide

1. Understanding the eBook Sol Chemistry Study Guide
  - The Rise of Digital Reading Sol Chemistry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Sol Chemistry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Sol Chemistry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Sol Chemistry Study Guide
  - Personalized Recommendations
  - Sol Chemistry Study Guide User Reviews and Ratings
  - Sol Chemistry Study Guide and Bestseller Lists
5. Accessing Sol Chemistry Study Guide Free and Paid eBooks
  - Sol Chemistry Study Guide Public Domain eBooks
  - Sol Chemistry Study Guide eBook Subscription Services
  - Sol Chemistry Study Guide Budget-Friendly Options
6. Navigating Sol Chemistry Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Sol Chemistry Study Guide Compatibility with Devices
  - Sol Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Sol Chemistry Study Guide
  - Highlighting and Note-Taking Sol Chemistry Study Guide
  - Interactive Elements Sol Chemistry Study Guide
8. Staying Engaged with Sol Chemistry Study Guide

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Sol Chemistry Study Guide
9. Balancing eBooks and Physical Books Sol Chemistry Study Guide
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Sol Chemistry Study Guide
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Sol Chemistry Study Guide
- Setting Reading Goals Sol Chemistry Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Sol Chemistry Study Guide
- Fact-Checking eBook Content of Sol Chemistry Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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