

**Worksheet – ASSIGNING OXIDATION NUMBERS****Oxidation Number Rules:**

1. The oxidation number of any pure element is 0.
2. The oxidation number of a monatomic ion equals that charge on the ion.
3. The more electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.
4. The oxidation number of fluorine in a compound is always -1.
5. Oxygen has an oxidation number of -2 unless it is combined with F, in which it is +1 or +2, or it is in peroxide (such as  $\text{H}_2\text{O}_2$  or  $\text{Na}_2\text{O}_2$ ), in which it is -1.
6. Hydrogen is +1, unless combined with a metal, and then it is -1.
7. In compounds, Group 1 is +1, Group 2 is +2, and Aluminum is +3.
8. The sum of the oxidation numbers of all atoms in a neutral compound is 0.
9. The sum of the oxidation numbers in a polyatomic ion equals the charge of the ion.

**Part A:** In the following questions, give the oxidation number of the indicated atoms/ion:

- |   |  |
|---|--|
| 1. N in $\text{N}_2\text{O}_5$ _____        | 16. C in $\text{CH}_4$ _____               |
| 2. S in $\text{H}_2\text{SO}_4$ _____       | 17. Mn in $\text{MnO}_2$ _____             |
| 3. C _____                                  | 18. S in $\text{SO}_3^{2-}$ _____          |
| 4. C in CO _____                            | 19. $\text{Mg}^{2+}$ _____                 |
| 5. Na in NaCl _____                         | 20. Cl _____                               |
| 6. H in $\text{H}_2\text{O}$ _____          | 21. $\text{O}_2$ _____                     |
| 7. Ba in $\text{BaCl}_2$ _____              | 22. P _____                                |
| 8. N in $\text{NO}_2^-$ _____               | 23. Na in $\text{Na}_2\text{S}$ _____      |
| 9. S in $\text{Al}_2\text{S}_3$ _____       | 24. S in $\text{H}_2\text{S}$ _____        |
| 10. S in $\text{HSO}_4^-$ _____             | 25. $\text{Ca}^{2+}$ _____                 |
| 11. Cl in $\text{Fe}(\text{ClO}_2)_3$ _____ | 26. C in $\text{CN}^-$ _____               |
| 12. Fe in $\text{Fe}(\text{ClO}_2)_3$ _____ | 27. H in $\text{OH}^-$ _____               |
| 13. N in $\text{NO}_3^-$ _____              | 28. Mn in $\text{KMnO}_4$ _____            |
| 14. $\text{Cu}^{2+}$ _____                  | 29. I in $\text{Mg}(\text{IO}_3)_2$ _____  |
| 15. $\text{Zn}^{2+}$ _____                  | 30. C in $\text{C}_2\text{O}_4^{2-}$ _____ |

# Oxidation Number Practice Answer Key

**Disha Experts**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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