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Oxidation = Gain **O** bonds, Lose **H** bonds



Strong ox agents

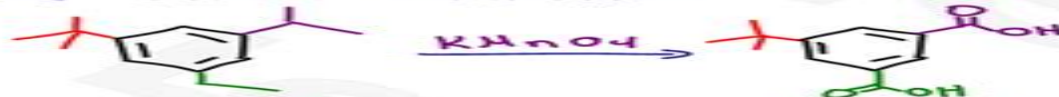
$\frac{CrO_3}{H_2O} \rightarrow H_2CrO_4$	$\frac{K_2Cr_2O_7}{H_2SO_4} \rightarrow K_2H_2O_4$
$\frac{Na_2Cr_2O_7}{H_2SO_4}$	



Alkene Oxidation (also alkynes)



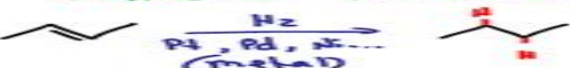
Side Chain Oxidation



only 1° 2° + 3°
get oxidized
4° = N/A

Reduction = Gain 4 bonds, Lose 0 bonds

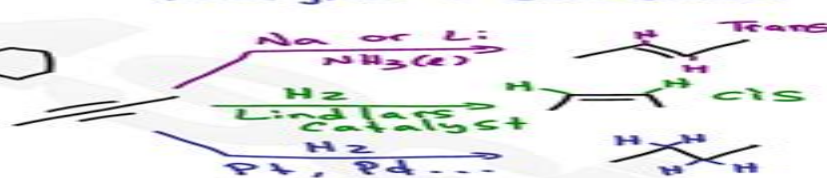
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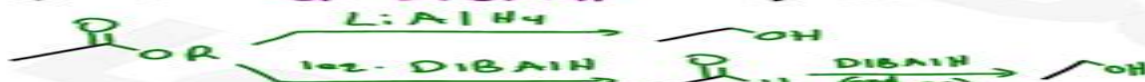
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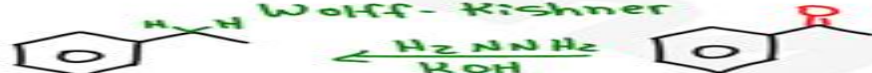
Alkyne Reduction



Reduction With Metal Hydrides



Wolff - Kishner



Clemmenson



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Oxidation And Reduction Study Guide

J Elliott



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