

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium *note: no d's, so the p. table referenced has Ac as 5f*
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (take a look at "4d"; must be 3 d not 4!)
- 17) $1s^2 2s^2 2p^6 3s^3 3d^5$ not valid (3p comes after 3s); also no 3 electrons in s
- 18) $[Ra] 7s^2 5f^8$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^3$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

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Table of Contents Small Scale Invest Electron Configuration Answers

1. Understanding the eBook Small Scale Invest Electron Configuration Answers
 - The Rise of Digital Reading Small Scale Invest Electron Configuration Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Small Scale Invest Electron Configuration Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Small Scale Invest Electron Configuration Answers
 - Personalized Recommendations
 - Small Scale Invest Electron Configuration Answers User Reviews and Ratings
 - Small Scale Invest Electron Configuration Answers and Bestseller Lists

5. Accessing Small Scale Invest Electron Configuration Answers Free and Paid eBooks
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 - Small Scale Invest Electron Configuration Answers eBook Subscription Services
 - Small Scale Invest Electron Configuration Answers Budget-Friendly Options
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 - Small Scale Invest Electron Configuration Answers Compatibility with Devices
 - Small Scale Invest Electron Configuration Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Small Scale Invest Electron Configuration Answers
 - Highlighting and Note-Taking Small Scale Invest Electron Configuration Answers
 - Interactive Elements Small Scale Invest Electron Configuration Answers
8. Staying Engaged with Small Scale Invest Electron Configuration Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Small Scale Invest Electron Configuration Answers
9. Balancing eBooks and Physical Books Small Scale Invest Electron Configuration Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Small Scale Invest Electron Configuration Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Small Scale Invest Electron Configuration Answers
 - Setting Reading Goals Small Scale Invest Electron Configuration Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Small Scale Invest Electron Configuration Answers
 - Fact-Checking eBook Content of Small Scale Invest Electron Configuration Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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