

**QUESTION 3 (Start on a new page)**

Compounds **A** to **E**, indicated in the table below, are used during two investigations to determine the factors which influence boiling point.

Investigation	Compound		Molecular mass (g·mol <sup>-1</sup> )	Boiling point (°C)
I	<b>A</b>	2,2-dimethyl propane	72	9
	<b>B</b>	2-methyl butane	72	27
	<b>C</b>	pentane	72	36
II	<b>D</b>	CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> OH	74	117
	<b>E</b>	CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CHO	72	75

3.1 Compounds **A**, **B** and **C** are structural isomers. Write down the:

3.1.1 Definition of the term *structural isomer* (2)

3.1.2 GENERAL FORMULA of the homologous series to which these compounds belong (1)

3.1.2 Type of structural isomerism illustrated by these compounds (1)

3.2 Consider the boiling points of the compounds in investigation I.

3.2.1 Define the term *boiling point*. (2)

3.2.2 Write down the independent variable for this investigation. (1)

3.2.3 Write down one control variable for this investigation. (1)

3.2.4 Explain fully why the boiling points increase from compound **A** to compound **C**. (3)

3.2.5 Which one of compounds **A** or **C** will have the highest vapour pressure at a certain temperature?  
Refer to the data in the table and give a reason for the answer. (2)

3.3 To which homologous series does compound **E** belong? (1)

3.4 Consider investigation II. Refer to the type of Van Der Waals forces in each of the compounds and explain why the boiling point of compound **D** is higher than that of compound **E**. (3)

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