

## NCERT Solutions for Class 11

### Chemistry

#### Chapter – 1 Some basic concepts of Chemistry

1. Calculate the molecular mass of the following:

i.  $\text{H}_2\text{O}$

Ans: The molecular mass of water,  $\text{H}_2\text{O}$ , can be calculated by the following steps given below:

= (2  $\times$  Atomic mass of hydrogen element) + (1  $\times$  Atomic mass of oxygen element)

= [(2  $\times$  1.0084 u) + (1  $\times$  16.00 u)]

= 2.016 u + 16.00 u

= 18.016 u or 18.02 u

ii.  $\text{CO}_2$

Ans: The molecular mass of carbon dioxide,  $\text{CO}_2$ , is calculated below:

= (1  $\times$  Atomic mass of carbon) + (2  $\times$  Atomic mass of oxygen)

= [(1  $\times$  12.011 u) + (2  $\times$  16.00 u)]

= 12.011 u + 32.00 u

= 44.01 u

iii.  $\text{CH}_4$

Ans: The molecular mass of methane,  $\text{CH}_4$ , is calculated below in step by step manner:

= (1  $\times$  Atomic mass of carbon) + (4  $\times$  Atomic mass of hydrogen)

= [(1  $\times$  12.011 u) + (4  $\times$  1.008 u)]

= 12.011 u + 4.032 u

= 16.043 u

2. Calculate the mass percent of different elements present in sodium sulphate ( $\text{Na}_2\text{SO}_4$ ).

Ans: The given compound in the question is sodium sulphate and its formula is  $\text{Na}_2\text{SO}_4$ . Its molecular formula is calculated below:

# Ncert Solutions For Class 11 Chemistry Chapter 1

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