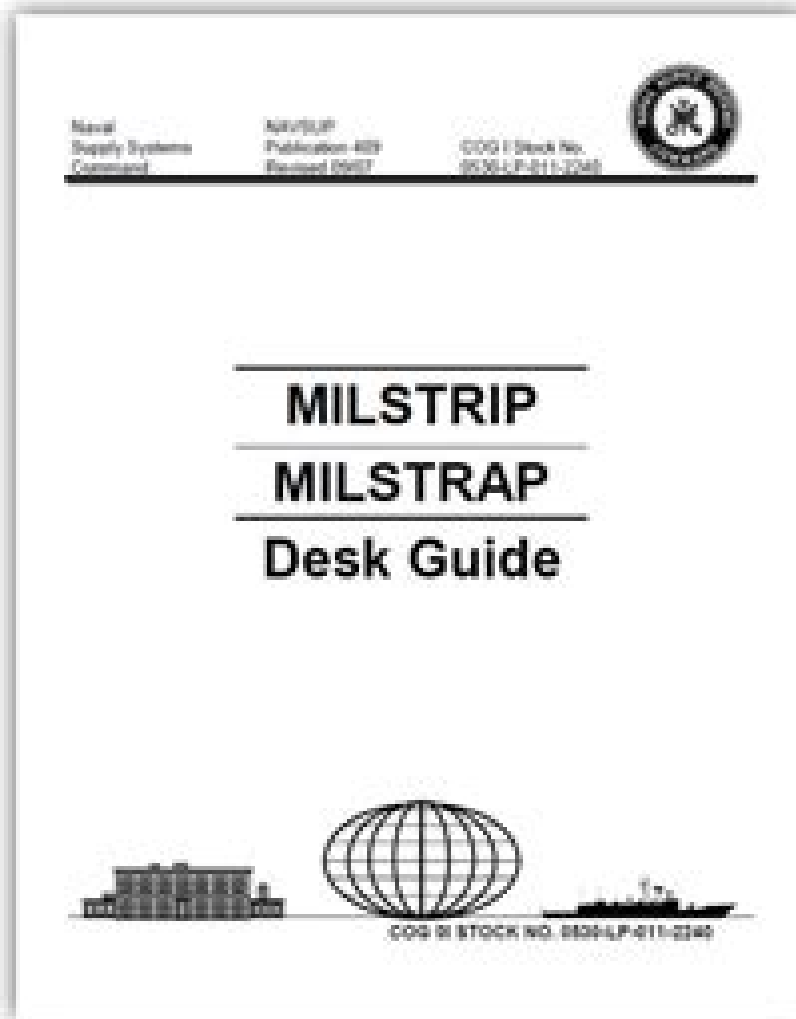




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web may 8 2017 after a time a reversible reaction in a closed system can reach what we call a dynamic equilibrium this is where the forwards reaction reactants reacting to produce the products and the backwards reaction products reacting to reform the reactants are occurring at the same rate

web all reactions tend towards a state of chemical equilibrium the point at which both the forward process and the reverse process are taking place at the same rate since the forward and reverse rates are equal the concentrations of the reactants and products are constant at equilibrium

web in a chemical equilibrium the forward and reverse reactions occur at equal rates and the concentrations of products and reactants remain constant law of mass action the law of mass action links the rate of a chemical reaction as proportional to the concentrations of the reactants and products in a chemical reaction for a chemical reaction

At equilibrium the rate of the forward reaction and the rate of the reverse reaction are equal. At equilibrium, both reactants and products are present. Some reaction mixtures contain mostly reactants and form only a few products at equilibrium. Some reaction mixtures contain mostly products and remain mostly reactants at equilibrium.

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$$\frac{1}{n} \sum_{i=1}^n x_i = \bar{x}$$
$$s_x = \sqrt{\frac{1}{n-1} \sum_{i=1}^n (x_i - \bar{x})^2}$$
$$y' = a + bx$$
$$r = \frac{s_y}{s_x} b$$

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