

Question 1. Calculate the molecular mass of the following:

(i) H_2O (ii) CO_2 (iii) CH_4

Answer: (i) Molecular mass of $\text{H}_2\text{O} = 2(1.008 \text{ amu}) + 16.00 \text{ amu} = 18.016 \text{ amu}$

(ii) Molecular mass of $\text{CO}_2 = 12.01 \text{ amu} + 2 \times 16.00 \text{ amu} = 44.01 \text{ amu}$

(iii) Molecular mass of $\text{CH}_4 = 12.01 \text{ amu} + 4 (1.008 \text{ amu}) = 16.042 \text{ amu}$

Question 2. Calculate the mass percent of different elements present in sodium sulphate ($\text{Na}_2 \text{SO}_4$).

Answer:

$$\text{Mass \% of an element} = \frac{\text{Mass of that element in the compound}}{\text{Molar mass of the compound}} \times 100$$

Now, $\text{Molar mass of } \text{Na}_2\text{SO}_4 = 2 (23.0) + 32.0 + 4 \times 16.0$
 $= 142 \text{ g mol}^{-1},$

$$\begin{aligned}\text{Mass percent of sodium} &= \frac{46}{142} \times 100 \\ &= 32.39 \%\end{aligned}$$

$$\begin{aligned}\text{Mass percent of sulphur} &= \frac{32}{142} \times 100 \\ &= 22.54 \%\end{aligned}$$

$$\begin{aligned}\text{Mass percent of oxygen} &= \frac{64}{142} \times 100 \\ &= 45.07 \%\end{aligned}$$

Question 3. Determine the empirical formula of an oxide of Iron which has 69.9 % iron and 30.1 % dioxygen by mass.

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VM Jensen



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Table of Contents Ncert Solution Chemistry 11 Chapter Answer

1. Understanding the eBook Ncert Solution Chemistry 11 Chapter Answer
 - The Rise of Digital Reading Ncert Solution Chemistry 11 Chapter Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Ncert Solution Chemistry 11 Chapter Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ncert Solution Chemistry 11 Chapter Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ncert Solution Chemistry 11 Chapter Answer

- Personalized Recommendations
 - Ncert Solution Chemistry 11 Chapter Answer User Reviews and Ratings
 - Ncert Solution Chemistry 11 Chapter Answer and Bestseller Lists
5. Accessing Ncert Solution Chemistry 11 Chapter Answer Free and Paid eBooks
- Ncert Solution Chemistry 11 Chapter Answer Public Domain eBooks
 - Ncert Solution Chemistry 11 Chapter Answer eBook Subscription Services
 - Ncert Solution Chemistry 11 Chapter Answer Budget-Friendly Options
6. Navigating Ncert Solution Chemistry 11 Chapter Answer eBook Formats
- ePub, PDF, MOBI, and More
 - Ncert Solution Chemistry 11 Chapter Answer Compatibility with Devices
 - Ncert Solution Chemistry 11 Chapter Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Ncert Solution Chemistry 11 Chapter Answer
 - Highlighting and Note-Taking Ncert Solution Chemistry 11 Chapter Answer
 - Interactive Elements Ncert Solution Chemistry 11 Chapter Answer
8. Staying Engaged with Ncert Solution Chemistry 11 Chapter Answer
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ncert Solution Chemistry 11 Chapter Answer
9. Balancing eBooks and Physical Books Ncert Solution Chemistry 11 Chapter Answer
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ncert Solution Chemistry 11 Chapter Answer
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ncert Solution Chemistry 11 Chapter Answer
- Setting Reading Goals Ncert Solution Chemistry 11 Chapter Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ncert Solution Chemistry 11 Chapter Answer

- Fact-Checking eBook Content of Ncert Solution Chemistry 11 Chapter Answer
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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