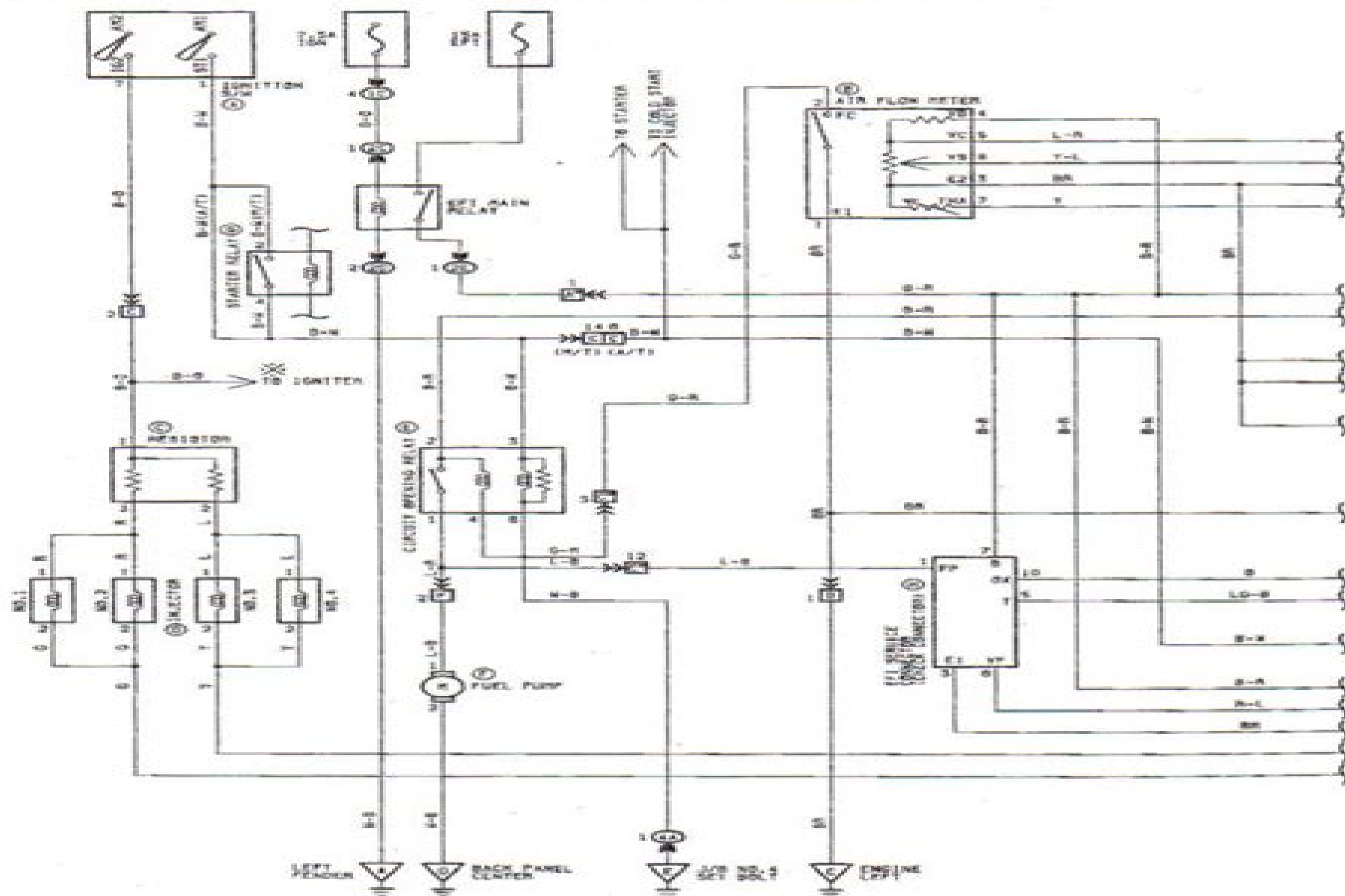


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web answer in a chemical reaction chemical equilibrium is the state in which the forward reaction rate and the reverse reaction rate are equal the result of this equilibrium is that the concentrations of the reactants and the products do not change

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web if the rate of the forward reaction decreases increases stays the same more products are formed in a chemical reaction at equilibrium the rate of the forward reaction is equal to the rate of the reverse reaction if the rate of the forward reaction increases more products are formed

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web in order to physically measure the rate of consumption of the bisulfite ion we will look for a color change in the reaction vessel the solution will change from clear to dark blue if you have the following equilibrium reaction $\text{CO} + 2\text{H}_2 \rightleftharpoons \text{CH}_3\text{OH}$

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web the energy needed for a reaction to occur the minimum amount of energy needed for a reaction to potentially occur the

energy added by a catalyst the energy possessed by the products question 5 60 seconds q increase in temperature of the reactants can do one of the following answer choices

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