

■ Review | Constants | Periodic Table

To determine whether a shiny gold-colored rock is actually gold, a chemistry student decides to measure its heat capacity. She first weighs the rock and finds it has a mass of 4.4 g. She then finds that upon absorption of 55.6 J of heat, the temperature of the rock rises from 25°C to 34°C .

Substance	Specific heat capacity, C_s ($\text{J}/(\text{g}\cdot^\circ\text{C})$)
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Elements

Lead 0.128

Gold 0.129

Silver 0.235

Copper 0.385

Iron 0.449

Aluminum 0.903

Compounds

Ethanol 2.42

Water 4.18

Materials

Glass (Pyrex) 0.75

You may want to reference (Pages 210 - 214) Section 6.4 while completing this problem.

Part A

Find the specific heat capacity of the substance composing the rock.

Express the specific heat in joules per gram-Celsius to two significant figures.

$C_s =$ $\text{J}/\text{g}\cdot^\circ\text{C}$

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Part B

Determine whether the value is consistent with the rock being pure gold by referring to the table.

- Yes, the value is consistent with the rock being pure gold.
- No, the value is inconsistent with the rock being pure gold.

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