

THERMOCHEMISTRY CALCULATIONS WORKSHEET 1

Using reaction equation ratios

- How much heat will be released when 6.44 g of Sulfur reacts with excess O_2 according to the following equation?
 $2S + 3O_2 \rightarrow 2SO_2 \quad \Delta H^\circ = -791.4 \text{ kJ}$
- How much heat will be released when 4.72 g of Carbon reacts with excess O_2 according to the following equation?
 $C + O_2 \rightarrow CO_2 \quad \Delta H^\circ = -393.5 \text{ kJ}$
- How much heat will be absorbed when 38.2 g of Bromine reacts with excess H_2 according to the following equation?
 $H_2 + Br_2 \rightarrow 2HBr \quad \Delta H^\circ = 72.80 \text{ kJ}$
- How much heat will be released when 1.48 g of Chlorine reacts with excess phosphorus according to the following equation?
 $2P + 3Cl_2 \rightarrow 2PCl_3 \quad \Delta H^\circ = -886 \text{ kJ}$
- How much heat will be released when 4.77 g of ethanol (C_2H_5OH) reacts with excess O_2 according to the following equation?
 $C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O \quad \Delta H^\circ = -1366.7 \text{ kJ}$
- How much heat will be absorbed when 13.7 g of Nitrogen reacts with excess O_2 according to the following equation?
 $N_2 + O_2 \rightarrow 2NO \quad \Delta H^\circ = -180 \text{ kJ}$
- How much heat will be released when 11.8 g of Iron reacts with excess O_2 according to the following equation?
 $3Fe + 2O_2 \rightarrow Fe_2O_3 \quad \Delta H^\circ = -1120.5 \text{ kJ}$
- How much heat will be released when 18.6 g of Hydrogen reacts with excess O_2 according to the following equation?
 $2H_2 + O_2 \rightarrow H_2O \quad \Delta H^\circ = -571.6 \text{ kJ}$
- How much heat will be transferred when 14.9 g of ammonia reacts with excess O_2 according to the following equation?
 $4NH_3 + 5O_2 \rightarrow 4NO + 6H_2O \quad \Delta H^\circ = -1170 \text{ kJ}$
- How much heat will be transferred when 5.81 g of graphite reacts with excess H_2 according to the following equation?
 $6C(\text{graphite}) + 3H_2 \rightarrow C_6H_6 \quad \Delta H^\circ = 49.03 \text{ kJ}$

Using $\Delta H = mC\Delta T$

- How many kilojoules of heat energy are required to heat all the aluminum (C_p of Al = $.902 \text{ J/g}^\circ\text{C}$) in a roll of aluminum foil (500.0 g) from room temperature (25.0°C) to the temperature of a hot oven (250.0°C)?
- One way to cool down your cup of coffee is to plunge an ice-cold piece of aluminum into it. Suppose you store a 20.0 g piece of aluminum (C_p of Al = $.902 \text{ J/g}^\circ\text{C}$) in the refrigerator at 4.40°C and then drop it into your coffee. The coffee temperature drops from 90.0°C to 55.0°C . How many joules of heat energy did the aluminum block absorb? (Ignore the cooling of the cup)
- Suppose you pick up a 16.0-kilogram ball of iron (such as a "shot-put" ball at a track event). The iron (C_p = $.451 \text{ J/g}^\circ\text{C}$) ball has the same temperature as the atmosphere on a cool day say 16.0°C . How many kilojoules of heat energy must the iron ball absorb to reach the temperature of your body (37.0°C)?
- The specific heat of silver is $0.24 \text{ J/g}^\circ\text{C}$. How much heat in joules must be added to a silver block of mass 86.0 g to raise its temperature by 9.00°C ?
- An 18.7 g sample of platinum metal increases in temperature by 2.30°C when 5.70 J of heat are added. What is the specific heat of platinum?

Answers: 1) -79.4kJ 2) -155kJ 3) 17.4kJ 4) -3.75kJ 5) -142kJ 6) -88.1kJ 7) -78.7kJ 8) -2660kJ 9) -256kJ 10) 3.96kJ
 11) 101kJ 12) 912J 13) 152kJ 14) 186J 15) $.132 \text{ J/g}^\circ\text{C}$

Thermochemistry Calculations Answer

Thure Georg Sahama, D. R. Torgeson



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

Thermochemical Calculations Ralph Richter Wenner, 1941 **Conceptual Chemistry Volume I For Class XI** S.K. Jain & Shailesh K. Jain, 1998 **Conceptual Chemistry Volume I For Class XI**

Thermochemistry Praveen Tyagi, 2006 Contents: Thermochemistry, The Specific Heat of Solids and Gases, Non Equilibrium and Classical Thermodynamics, Kinetic Molecular Theory of Gases, Calculation of Phase Diagrams and Thermochemistry of Alloy Phases Y. Austin Chang, John F. Smith, 1979 **Physico-chemical Calculations** Joseph Knox, 1919 *A Text Book of Thermo-chemistry and Thermodynamics* Otto Sackur, 1917 *The Encyclopaedia Britannica*, 1911 *Thermodynamic Data on Oxides and Silicates* Surendra K. Saxena, Nilanjan Chatterjee, Yingwei Fei, Guoyin Shen, 2012-12-06 During the last thirty years, profound developments in experimental techniques to measure high temperature and pressures and thermodynamic properties of minerals have occurred. This technical development has been matched by an increased sophistication in applying theoretical methods to obtain new data or improve the quality of existing data. Using these new techniques, *Assessed Thermodynamic Data on Oxides and Silicates* represents the successful attempt of the authors to develop an internally systematized data base which satisfies the constraints of calorimetric measurements, phase equilibrium data, measured thermophysical properties of a phase, and heat capacities and entropies estimated from lattice vibrational models.

Encyclopaedia Britannica: Submarine Mines-Tom-Tom Hugh Chisholm, 1911 This eleventh edition was developed during the encyclopaedia's transition from a British to an American publication. Some of its articles were written by the best known scholars of the time, and it is considered to be a landmark encyclopaedia for scholarship and literary style.

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