

# Chemistry

## Thermochemistry

### Lesson 10

#### Lesson Plan

David V. Fansler

#### The Flow of Energy-Heat

*Objectives: Explain the relationship between energy and heat; Distinguish between heat capacity and specific heat.*

- Energy and Heat
  - o **Thermochemistry** is concerned with the heat changes that occur during chemical reactions.
  - o **Energy** is the capacity for doing work or supplying heat.
    - **Work** is a force times a distance.
    - Energy is not visible – it is odorless and tasteless. Energy can only be detected because of its results.
  - o Energy stored within the chemical structure units of chemical substances is called **chemical potential energy**.
    - Gasoline is a form of stored potential energy, so is sugar.
    - The amount of energy stored in a chemical depends on the elements and their arrangements.
  - o **Heat**, represented by  $q$ , is energy that transfers from one object to another due to a temperature difference between them.
    - As with energy, heat itself cannot be detected, only the affects of adding or subtracting heat can be detected.
    - An effect of adding heat is a rise in temperature of an object.
    - Heat always flows from a warmer object to a cooler object.
- Exothermic and Endothermic Processes
  - o Almost all chemical reactions and changes in state involve the release or absorption of heat.
    - When looking at the process, we define the **system** as those objects you are focusing on.
    - The **surroundings** are everything else.
    - Together, the system and the surroundings make up the **universe**.
  - o A major study of Thermochemistry is to examine the flow of heat from a system to the surroundings or from the surroundings to the system
  - o The **Law of Conservation of Energy** states that in any chemical or physical process that energy is neither created nor destroyed.
    - All energy in a process can be accounted for as work, stored energy or heat.
  - o When a reaction releases heat into it's surrounding, it is termed an **exothermic process** –  $q$  is defined as negative.
  - o When a reaction absorbs heat from its surroundings, it is termed an **endothermic process** –  $q$  is defined as positive.
  - o Examples

# Thermochemistry Lesson Ideas

**Peipei Pang**



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