Thermochemistry Multiple Choice Questions

	1	When the three types of energy changes are arranged in order of decreasing molar heat values, what is the correct order? A. chemical, phase, nuclear B. nuclear, chemical, phase D. phase, chemical, nuclear			
	2.	The temperature remains constant as energy is added to a substance. How may the substance be changing? A. from a gas to a solid B. from a liquid to a gas D. in the amount of kinetic energy.			
A	3.	 Which metal requires the most energy to raise 1.00 g of it b. A. aluminum. B. copper C. 	y 1.00°C? lead	D. nickel	
•	4	A substance has a molar heat of combustion of -810.4 kJ calorimeter containing 8.60 kg of water, what is the increase A. 0.156°C B. 6.41°C C.			
	5.	The addition of 9.54 k J of heat is required to raise the ten to 45.0°C. What is the heat capacity of this hydrocarbon? A. 0.94 J/g C B. 1.73 J/g °C C.	mperature of 225.0 g of a 188 J/g≠C	liquid hydrocorbon from 20.5°C D. 9.42 J/g*C	
C	6.	. If the molar heat of fusion of sodium is 2.63 k.J/mol, how melting point? A. 2.63 k.J. B. 2.93 k.J. C.		elt 180.0 g of solid sodium at its	
	7.		the sublimation of dry ice the vaporization of water		
A	8.	If the heat of formation for the chemical reaction, S ₁ (x) = equation be rewritten to include the heat term? A S ₂ (x) + 12O ₂ (y) → 85O ₂ (y) + 3166 kJ C. B. S ₂ (x) + 12O ₂ (y) → 85O ₂ (y) - 395.7 kJ D.			
A	9.	Which statement correctly describes the energy changes that take place when a solid changes to a liquid at constant temperature? A. The potential energy increases, and the kinetic energy remains constant. B. The potential energy remains constant, and the kinetic energy increases. C. The potential energy decreases, and the kinetic energy decreases. D. The potential energy increases, and the kinetic energy decreases.			
	20.	Why does warm water have a more rapid note of evaporation than cold water? A. It has a higher viscosity. B. Its molecules have a higher overage kinetic energy. C. Its molecules have more attraction for one another. D. Alere of its molecules have morely the same kinetic energy.			
	11.	If the heat of fusion of a substance is 20 kJ/mol , what heat is released when 1.0 mol of liquid at the melting point freezes? A. 10 kJ B. 20 kJ C. 40 kJ D. 80 kJ			
c	12.	The amount of 2.00 mol of a solid ionic compound was dissolved in 225 mL of water in an insulated container. The initial temperature of the water was 47.8°C. What is Hoolin? A27.8 k J/mol D6.26 k J/mol D6.26 k J/mol			
b	13	3. What does it mean if the AH value for a chemical reaction is A. kinetic energy is increasing in the system B. potential energy is decreasing in the system C. products have less potential energy than reactionts D. reactionts have less potential energy than products.	i patential energy is decreasing in the system products have less potential energy than reactants		
A	34.	 Living plants produce glucose in the process of photosynthes ^{6CO₂(g) + 6H₂O(f) = energy → C₄S Is this reaction endothermic or exothermic, and is the value A endothermic, positive B endothermic, negative C.} 	luO _u (x) + 6O _x (g) of ΔH* positive or negativ		

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Internal Combustion Engines Shyam K. Agrawal, 2006 Salient Features The New Edition Is A Thoroughly Revised Version Of The Earlier Edition And Presents A Detailed Exposition Of The Basic Principles Of Design Operation And Characteristics Of Reciprocating I C Engines And Gas Turbines Chemistry Of Combustion Engine Cooling And Lubrication Requirements Liquid And Gaseous Fuels For Ic Engines Compressors Supercharging And Exhaust Emission Its Standards And Control Thoroughly Explained Jet And Rocket Propulsion Alternate Potential Engines Including Hybrid Electric And Fuel Cell Vehicles Are Discussed In Detail Chapter On Ignition System Includes Electronic Injection Systems For Si And Ci Engines 150 Worked Out Examples Illustrate The Basic Concepts And Self Explanatory Diagrams Are Provided Throughout The Text More Than 200 Multiple Choice Questions With Answers A Good Number Of Review Questions Numerical With Answers For Practice Will Help Users In Preparing For Different Competitive Examinations With These Features The Present Text Is Going To Be An Invaluable One For Undergraduate Mechanical Engineering Students And Amie Candidates

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