

## PRACTICE TEST/THERMOCHEMISTRY/AP CHEMISTRY

1. a) Calculate the amount of heat transferred when 10.00 g of  $\text{N}_2\text{O}(\text{g})$  is formed by the following reaction:



- b) Draw an energy diagram for this process.

2. Predict the value for  $\Delta H^\circ_f$  for the following scenarios:

- a)  $\text{Br}_2(\text{g})$
- b)  $\text{Br}_2(\text{l})$
- c)  $\text{I}_2(\text{g})$
- d)  $\text{I}_2(\text{s})$

3. Calculate the  $\Delta H_{\text{rxn}}$  for the following reaction:



$$\begin{aligned}\Delta H^\circ_f \text{ C}_2\text{H}_4(\text{g}) &= 226.6 \text{ kJ/mol} \\ \Delta H^\circ_f \text{ CO}_2(\text{g}) &= -393.5 \text{ kJ/mol} \\ \Delta H^\circ_f \text{ H}_2\text{O}(\text{l}) &= -285.8 \text{ kJ/mol}\end{aligned}$$

4. A 5.00 g sample of liquid water at  $25.0^\circ\text{C}$  is heated by the addition of 84.0 J of energy. The final temperature of the water is \_\_\_\_\_  $^\circ\text{C}$ . The specific heat capacity of the liquid is  $4.18 \text{ J/g}^\circ\text{C}$ .

- a) -21      b) 4.02      c) 25.2      d) 95.2      e) 29.0

5. *THIS PROBLEM WAS TAKEN DIRECTLY FROM THE 1995 AP CHEM. EXAM.*  
Propane is a hydrocarbon that is commonly used as a fuel for cooking.  
Propane's formula is  $\text{C}_3\text{H}_8$ .

- a) Write a balanced equation for the complete combustion of propane gas.
- b) Calculate the volume of air at  $30^\circ\text{C}$  and 1.00 atm that is needed to burn completely 10.0 g of propane. Assume that air is 21.0%  $\text{O}_2$  by volume.
- c) The heat of combustion ( $\Delta H^\circ_{\text{combustion}}$ ) is  $-2,220.1 \text{ kJ/mol}$ . Calculate the heat of formation,  $\Delta H^\circ_f$ , of propane given that  $\Delta H_f$  of  $\text{H}_2\text{O}(\text{l})$  is  $-285.3 \text{ kJ/mol}$  and  $\Delta H^\circ_f$  of  $\text{CO}_2(\text{g})$  is  $-393.5 \text{ kJ/mol}$ .
- d) Assuming that all of the heat evolved burning 10.0 g propane is transferred to 8.00 kg of water (specific heat =  $4.184 \text{ J/g}^\circ\text{C}$ ), calculate the increase in temperature of the water.

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