

Thermochemistry and Gas Laws Unit Review

Energy Math

1. How much energy is released when a brick with a mass of 329g is cooled from 80°C to 50°C? The specific heat capacity of brick is 0.84 J/g°C.

$$Q = mc\Delta T$$

$$(329g)(0.84J/g^{\circ}C)(-30^{\circ}C)$$

$$Q = -8290J$$

2. What is the mass of a hotdog that releases 3600 J of energy as it is cooled from 95°C to 45 °C? The specific heat capacity of a hotdog is 1.5 J/g°C.

$$m = \frac{-3600J}{(1.5J/g^{\circ}C)(-50^{\circ}C)} = 48g$$

3. What is the specific heat capacity of a 350 g object that rises in temperature 15°C when 55,900 J of energy are added?

$$C = \frac{55900J}{(350g)(15^{\circ}C)} = 10.6J/g^{\circ}C$$

4. 45,920 J of energy are added to a 365g block of copper originally at 35°C. What will the new temperature of the copper be? The specific heat capacity of copper is 0.385 J/g°C.

$$\Delta t = \frac{45920J}{(365g)(0.385J/g^{\circ}C)} = 326^{\circ}C = T_f - 35^{\circ}C$$

$$+35^{\circ}C \quad +35^{\circ}C$$

$$361^{\circ}C = T_f$$

5. 89,290 J of energy are removed from a 670 g block of aluminum originally at 125°C. What will the new temperature of the aluminum be? The specific heat capacity of aluminum is 0.902 J/g°C.

$$\Delta t = \frac{-89290J}{(670g)(0.902J/g^{\circ}C)} = -148^{\circ}C = T_f - 125^{\circ}C$$

$$+125^{\circ}C \quad +125^{\circ}C$$

$$-23^{\circ}C = T_f$$

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Ice Cream Review, 1927

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