

Thermochemistry Sample Test Questions

Identify the choice that best completes the statement or answers the question.

1. What happens to the energy produced by burning gasoline in a car engine?
A) The energy heats the parts of the engine.
B) The energy is transformed into work to move the car.
C) The energy is lost as heat in the exhaust.
D) all of the above
2. A piece of metal is heated, then submerged in cool water. Which statement below describes what happens?
A) The temperature of the water will increase and the temperature of the metal will decrease.
B) The temperature of the water will increase.
C) The temperature of the metal will increase.
D) The temperature of the water will decrease.
3. How does a calorie compare to a joule?
A) A calorie is smaller than a joule.
B) A calorie is equal to a joule.
C) The relationship cannot be determined.
D) A calorie is larger than a joule.
4. What would likely happen if you were to touch the flask in which an endothermic reaction were occurring?
A) The flask would feel the same as before the reaction started.
B) The flask would probably feel cooler than before the reaction started.
C) The flask would probably feel warmer than before the reaction started.
D) none of the above
5. Which of the following is NOT a form of energy?
A) heat
B) pressure
C) light
D) electricity
6. When energy is changed from one form to another, _____.
A) all of the energy is changed to a useful form
B) some of the energy is lost entirely
C) a physical change occurs
D) all of the energy can be accounted for
7. If heat is released by a chemical system, an equal amount of heat will be _____.
A) released by the surroundings
B) released by the universe
C) absorbed by the surroundings
D) absorbed by the universe
8. Which of the following is transferred due to a temperature difference?
A) mechanical energy
B) chemical energy
C) heat
D) electrical energy
9. In an exothermic reaction, the energy stored in the chemical bonds of the reactants is _____.
A) greater than the energy stored in the bonds of the products.
B) less than the energy stored in the bonds of the products.
C) less than the heat released
D) equal to the energy stored in the bonds of the products
10. A process that absorbs heat is a(n) _____.
A) ectothermic process
B) polythermic process
C) exothermic process
D) endothermic process
11. When your body breaks down sugar completely, how much heat is released compared to burning the same amount of sugar in a flame?
A) The body releases more heat.
B) The body releases no heat.
C) The body releases less heat.
D) The body releases the same amount of heat.
12. The quantity of heat required to change the temperature of 1 g of a substance by 1°C is defined as _____.
A) specific heat
B) a joule
C) a calorie
D) density

Thermochemistry Test Questions

M Woodhall



Thermochemistry Test Questions:

Reviewing **Thermochemistry Test Questions**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Thermochemistry Test Questions**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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