

Thermochemistry Review Part 1

Definition Questions: you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is $0.5050 \text{ J/g}^\circ\text{C}$. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of 50 g and an initial temperature of 90°C is placed into 100 mL of water at a temperature of 25°C . The temperature of water rises to 31.3°C . Determine the specific heat capacity of aluminum.

$Q \text{ lost (Al)} = Q \text{ gained (Water)}$

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

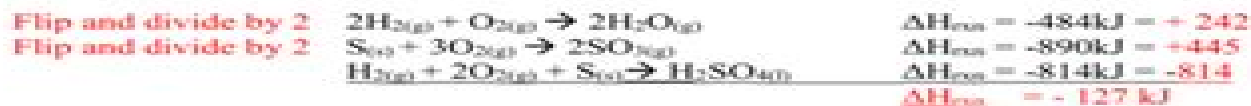
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by 300 g of methanol, CH_3OH , as it evaporates? Given: The molar heat of vaporization is 35.3 kJ/mol .

$q = ?$	$\Delta H = q/n$	$n \text{ CH}_3\text{OH} = m/\text{MM}$
$m = 300 \text{ g}$	$q = H \times n$	$n = 300 \text{ g} / 32 \text{ g/mol}$
$\Delta H = 35.3 \text{ kJ/mol}$	$q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$	$n = 9.375 \text{ mol}$
	$q = 332.81 \text{ kJ}$	

4. What is the molar enthalpy of the formation of $1 \text{ mol H}_2\text{SO}_{4(l)}$ given the following information?



Thermochemistry 1 Answers

Otto Sackur



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Table of Contents Thermochemistry 1 Answers

1. Understanding the eBook Thermochemistry 1 Answers
 - The Rise of Digital Reading Thermochemistry 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry 1 Answers
 - Personalized Recommendations

- Thermochemistry 1 Answers User Reviews and Ratings
- Thermochemistry 1 Answers and Bestseller Lists
- 5. Accessing Thermochemistry 1 Answers Free and Paid eBooks
 - Thermochemistry 1 Answers Public Domain eBooks
 - Thermochemistry 1 Answers eBook Subscription Services
 - Thermochemistry 1 Answers Budget-Friendly Options
- 6. Navigating Thermochemistry 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry 1 Answers Compatibility with Devices
 - Thermochemistry 1 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry 1 Answers
 - Highlighting and Note-Taking Thermochemistry 1 Answers
 - Interactive Elements Thermochemistry 1 Answers
- 8. Staying Engaged with Thermochemistry 1 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry 1 Answers
- 9. Balancing eBooks and Physical Books Thermochemistry 1 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry 1 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry 1 Answers
 - Setting Reading Goals Thermochemistry 1 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry 1 Answers
 - Fact-Checking eBook Content of Thermochemistry 1 Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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