

## Thermochemistry (Energy changes in chemical reactions)

Energy = capacity to do work

$W = F \times d$  something that an object possesses exchanged through heat + work

Kinetic energy = energy of motion

$$KE = \frac{1}{2}mv^2 = J$$

Potential energy = energy of position, composition, condition

Thermal energy = energy associated with random molecular motion (KE)

↳ chemical energy

↳ measure as temperature

Thermal Energy or T

### Review of important terms in thermochemistry:

- **System**-part of the universe under study chemical rxn / physical process
- **Surroundings**-part of the universe outside of the system solution / solvents, cups, air
- **Universe**-system + surroundings

Heat is a form in which energy can be transferred between a system and its surroundings caused by a temperature difference. Energy is transferred until the temperature of the two are equal.

Thermal equilibrium

The quantity of heat ( $q$ ) required to change the temperature of a substance depends on:

- depends on amount of substance / system
- identity of substance
- $\Delta T$

Note: the SI unit of energy is the **Joule (J)**. A Joule is equal to a  $\text{kg m}^2/\text{s}^2$ . Other common units of energy are: **kilojoules (kJ)**, **calories (cal)** where  $1 \text{ cal} = 4.184 \text{ J}$  and **kilowatt-hour (kWh)** where  $1 \text{ kWh} = 3.60 \times 10^6 \text{ J}$ .

Caloric = 1000 cal

We can measure the amount of thermal energy transfer ( $q$ ) using either of the equations below depending on the experimental setup (or what is provided in a problem):

$$q = mc\Delta T$$

$$q = C\Delta T$$

where  $m$  = mass (usually in grams) and  $\Delta T = T_f - T_i$

$c$  = specific heat capacity (usually  $\text{J/g}^\circ\text{C}$ ) = quantity of heat required to change the temp of 1 g of a substance by  $1^\circ\text{C}$

↳ intensive property

$C$  = heat capacity (usually  $\text{J/}^\circ\text{C}$ ) = quantity of heat required to change the temp of a system by  $1^\circ\text{C}$

↳ extensive property

- depends on amount present

# Thermochemistry Answers Energy Changes

**E. N. Ramsden**



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**Chemistry and Physics of Energetic Materials** S.N. Bulusu, 2012-12-06 This book represents a collection of lectures presented at the NATO Advanced study Institute ASI on Chemistry Physics of the Molecular Processes in Energetic Materials held at Hotel Torre Normanna Altavilla Milicia Sicily Italy September 3 to 15 1989 The institute was attended by seventy participants including twenty lecturers drawn from thirteen countries The purpose of the institute was to review the major advances made in recent years in the theoretical and experimental aspects of explosives and propellants In accordance with the format of the NATO ASI it was arranged to have a relatively small number of speakers to present in depth review type lectures emphasizing the basic research aspects of the subject over a two week period Most of the speakers gave two lectures each in excess of one hour with additional time for discussions The scope of the meeting was limited to molecular and spectroscopic studies since the hydrodynamic aspects of detonation and various performance criteria of energetic materials are often covered adequately in other international meetings An attempt was made to have a coherent presentation of various theoretical computational and spectroscopic approaches to help a better understanding of energetic materials from a molecular point of view The progress already made in these areas is such that structure property e.g.

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