

Thermochemistry Worksheet #1 KEY

1. The reaction of magnesium with sulfuric acid was carried out in a calorimeter. This reaction caused the temperature of 27.0 grams of liquid water, within the calorimeter, to raise from 25.0°C to 76.0°C. Calculate the energy associated with this reaction. **Answer: 5760 J**

$$q = s \times m \times \Delta T = (4.18 \text{ J/g}^\circ\text{C}) \times (27.0 \text{ g}) \times (51.0^\circ\text{C}) = 5755.86 \text{ J} = 5760 \text{ J}$$

2. The reaction of zinc with nitric acid was carried out in a calorimeter. This reaction caused the temperature of 72.0 grams of liquid water, within the calorimeter, to raise from 25.0°C to 100.°C. Calculate the energy associated with this reaction. **Answer: 22,600J**

$$q = s \times m \times \Delta T = (4.18 \text{ J/g}^\circ\text{C}) \times (72.0 \text{ g}) \times (75.0^\circ\text{C}) = 22572 \text{ J} = 22,600 \text{ J}$$

3. A 4.00 gram sample of solid gold was heated from 274K to 314K. How much energy was involved? Specific heat capacity of gold is 0.129 J/g°C) **Answer: 20.6J**

$$q = s \times m \times \Delta T = (0.129 \text{ J/g}^\circ\text{C}) \times (4.00 \text{ g}) \times (40\text{K}) = 20.64 \text{ J} = 20.6 \text{ J}$$

Use ΔH_f° values to solve the following problems.

4. Determine the ΔH for each of the following reactions. Classify each reaction as either exothermic or endothermic.



$$\Delta H_{\text{rxn}} = [4(-241.8 \text{ kJ/mol}) + 3(-393.5 \text{ kJ/mol})] - [1(-103.9 \text{ kJ/mol}) + 5(0 \text{ kJ/mol})] = -2043.8 \text{ kJ/mol}$$

exothermic



$$\Delta H_{\text{rxn}} = [2(-92.3 \text{ kJ/mol})] - [1(0 \text{ kJ/mol}) + 1(0 \text{ kJ/mol})] = -184.6 \text{ kJ/mol}$$

exothermic



$$\Delta H_{\text{rxn}} = [(-393.5 \text{ kJ/mol}) + 2(-241.8 \text{ kJ/mol})] - [1(-74.85 \text{ kJ/mol}) + 2(0 \text{ kJ/mol})] = 165.0 \text{ kJ/mol}$$

endothermic



Thermochemistry Calculations Answer Key

William L. Masterton, Emil J. Slowinski



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