

Thermochemistry Test Review: Multiple Choice Problems

- What is the correct order, from highest to lowest molar heat value, for the three types of energy change?
 - Chemical, phase, nuclear
 - Phase, chemical, nuclear
 - Nuclear, phase, chemical
 - Nuclear, chemical, phase
- Which metal requires the most energy to raise 1.00 g of it by 1.00°C? (Refer to Chemistry Data Booklet)
 - Aluminum
 - Copper
 - Lead
 - Nickel
- What is the specific heat capacity of a substance if 2.41×10^4 J are needed to change the temperature of 105.0 g of it from 25.0°C to 250.0°C?
 - $1.02 \times 10^{-4} \text{ J/g}^\circ\text{C}$
 - $9.18 \times 10^{-4} \text{ J/g}^\circ\text{C}$
 - $0.918 \text{ J/g}^\circ\text{C}$
 - $1.02 \text{ J/g}^\circ\text{C}$
- Which statement about enthalpy is true?
 - Heat is given off to the surroundings in endothermic reactions.
 - Some substances have a negative specific heat capacity.
 - Specific heat capacity is the same for all liquids.
 - The sign of ΔH is always negative in exothermic reactions.
- What happens to the value of ΔH for a thermochemical reaction if the reaction is reversed?
 - ΔH has the same numerical value, and the sign changes.
 - ΔH has the same numerical value, and the sign remains the same.
 - ΔH is the reciprocal of the original value, and the sign changes.
 - ΔH is the reciprocal of the original value, and the sign remains the same.
- Which is an exothermic process?
 - Ice melting
 - Water boiling
 - Water evaporating
 - Water vapour condensing
- Which statement is true for the combustion of ethanol?

$$\text{C}_2\text{H}_5\text{OH}(\text{l}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{l}) \quad \Delta H = -1370 \text{ kJ}$$
 - The enthalpy change would be the same if gaseous water were produced.
 - The potential energy of the products is less than the potential energy of the reactants.
 - The products of the reaction occupy a larger volume than the reactants.
 - The reaction is endothermic.
- What is the energy required to evaporate two moles of liquid water given the following equations?

$$\begin{aligned} 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g}) &\rightarrow 2\text{H}_2\text{O}(\text{g}) & \Delta H &= -483.6 \text{ kJ} \\ 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g}) &\rightarrow 2\text{H}_2\text{O}(\text{g}) & \Delta H &= -571.6 \text{ kJ} \end{aligned}$$
 - 44.0 kJ
 - 88.0 kJ
 - 527.6 kJ
 - 1055.2 kJ
- Which statement correctly describes an endothermic chemical reaction?
 - The products have higher potential energy than the reactants, and the ΔH is negative.
 - The products have higher potential energy than the reactants, and the ΔH is positive.
 - The products have lower potential energy than the reactants, and the ΔH is negative.
 - The products have lower potential energy than the reactants, and the ΔH is positive.
- Given the thermochemical equation,

$$2\text{NO}_2(\text{g}) \rightarrow 2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \quad \Delta H = +114 \text{ kJ}$$
 what is the ΔH for the reaction $\text{NO}(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$?
 - 114 kJ
 - 57 kJ
 - +57 kJ
 - +114 kJ
- Using the data below, what is the order of changes that occur when ethanol is heated from -25.0°C to 85.0°C?

Boiling point of ethanol: 78.3°C
Melting point of ethanol: -117.3°C

 - Phase change, temperature change
 - Phase change, temperature change, phase change
 - Temperature change, phase change
 - Temperature change, phase change, temperature change
- Which statement best describes the formation of methanol, $\text{CH}_3\text{OH}(\text{l})$, from the elements? (Refer to Data Booklet, page 5)
 - It is endothermic and heat is absorbed.
 - It is exothermic and heat is released.
 - It is exothermic and heat is absorbed.
 - It is exothermic and heat is released.
- Which process results in the greatest endothermic change for 10.0 g of H_2O ?
 - Condensation
 - Melting
 - Solidification
 - Vaporization
- The following decomposition reaction may occur in an air bag:

$$2\text{NaN}_3(\text{s}) \rightarrow 3\text{N}_2(\text{g}) + 2\text{Na}(\text{s}) \quad \Delta H = -43.5 \text{ kJ}$$
 What is the heat of formation, ΔH_f° , for NaN_3 ?
 - 43.5 kJ
 - 21.8 kJ
 - 21.8 kJ
 - 43.5 kJ
- A small sample released 2.0×10^{11} kJ of energy while undergoing a change. What type of change most likely occurred?
 - Chemical
 - molecular
 - nuclear
 - physical

Thermochemistry Multiple Choice Test

Judd E. Hollander



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