

Assignment 10: Fluoroborates

1. Draw the structural formulae of compounds (a)-(f):



Standard BF_3 and



2. A gas-phase ion-molecule study is carried out with BF_3 and gaseous ions. Identify which ion is usually of first order in collision experiments. Calculate the bond energy of that ion (standard $\Delta_{\text{f}}H^\circ$ for BF_3 is $-1036 \text{ kJ mol}^{-1}$, $\Delta_{\text{f}}H^\circ$ for BF_3^+ is $-1036 \text{ kJ mol}^{-1}$).

3. A BF_3 molecule at absolute zero, at 0 K , is vibrational ground state. It is in a p_x orbital. What is the bond energy of BF_3 ? What is the bond energy of BF_3^+ ?

What is $\Delta_{\text{f}}H^\circ$ for BF_3 ?

4. When a BF_3 molecule at 0 K is vibrational ground state, it is in a p_x orbital. What is the bond energy of BF_3 ? What is the bond energy of BF_3^+ ?

5. The vibrational frequency of BF_3 is 1036 cm^{-1} . What is the bond energy of BF_3 ? What is the bond energy of BF_3^+ ?

6. Draw the structural formulae of the following compounds. What is the bond energy of BF_3 ?



7. A molecule is formed by adding BF_3 to a gas-phase ion. It is in a p_x orbital. What is the bond energy of BF_3 ? What is the bond energy of BF_3^+ ?

8. By convention, the bond energy of BF_3 is 1036 kJ mol^{-1} . What is the bond energy of BF_3^+ ?

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Table of Contents Thermochemistry Practice Calculation

1. Understanding the eBook Thermochemistry Practice Calculation
 - The Rise of Digital Reading Thermochemistry Practice Calculation
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Practice Calculation
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Practice Calculation
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Practice Calculation
 - Personalized Recommendations
 - Thermochemistry Practice Calculation User Reviews and Ratings
 - Thermochemistry Practice Calculation and Bestseller Lists
5. Accessing Thermochemistry Practice Calculation Free and Paid eBooks
 - Thermochemistry Practice Calculation Public Domain eBooks
 - Thermochemistry Practice Calculation eBook Subscription Services
 - Thermochemistry Practice Calculation Budget-Friendly Options

6. Navigating Thermochemistry Practice Calculation eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Practice Calculation Compatibility with Devices
 - Thermochemistry Practice Calculation Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Practice Calculation
 - Highlighting and Note-Taking Thermochemistry Practice Calculation
 - Interactive Elements Thermochemistry Practice Calculation
8. Staying Engaged with Thermochemistry Practice Calculation
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Practice Calculation
9. Balancing eBooks and Physical Books Thermochemistry Practice Calculation
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Practice Calculation
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermochemistry Practice Calculation
 - Setting Reading Goals Thermochemistry Practice Calculation
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermochemistry Practice Calculation
 - Fact-Checking eBook Content of Thermochemistry Practice Calculation
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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