

Thermochemistry Practice Problems

KEY

Flow of Energy:

1. How many kJ of energy are in a donut that contains 200.0 Calories?

$$\frac{200 \text{ Cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ Cal}} = \boxed{836.8 \text{ kJ}}$$

2. What is the specific heat of a substance that has a mass of 25.0 g and requires 525.0 calories to raise its temperature by 15.0°C?

$$\begin{aligned} m &= 25.0 \text{ g} \\ q &= 525.0 \text{ cal} \\ \Delta T &= 15.0^\circ\text{C} \end{aligned}$$

$$q = mc\Delta T$$

$$525.0 \text{ cal} = (25.0 \text{ g})(c)(15.0^\circ\text{C})$$

$$\boxed{c = 1.40 \text{ cal/g}^\circ\text{C}}$$

3. Suppose 100.0 g of H₂O(s) absorbs 1255.0 J of heat. What is the corresponding temperature change? The specific heat capacity of H₂O(s) is 2.1 J/g°C.

$$\begin{aligned} q &= 1255.0 \text{ J} \\ m &= 100.0 \text{ g} \\ c &= 2.1 \text{ J/g}^\circ\text{C} \end{aligned}$$

$$q = mc\Delta T$$

$$1255.0 \text{ J} = (100.0 \text{ g})(2.1 \text{ J/g}^\circ\text{C})(\Delta T)$$

$$\boxed{\Delta T = 6.0^\circ\text{C}}$$

4. How many joules of heat energy are required to raise the temperature of 100.0 g of aluminum by 120.0°C? The specific heat capacity of aluminum is 0.90 J/g°C.

$$\begin{aligned} m &= 100.0 \text{ g} \\ c &= 0.90 \text{ J/g}^\circ\text{C} \\ \Delta T &= 120.0^\circ\text{C} \end{aligned}$$

$$q = (100.0 \text{ g})(0.90 \text{ J/g}^\circ\text{C})(120.0^\circ\text{C})$$

$$\boxed{q = 11,000 \text{ J}}$$

5. In reference to #4, what is the molar heat capacity of aluminum?

$$(0.90 \text{ J/g}^\circ\text{C})(26.98 \text{ g/mol}) = \boxed{24 \text{ J/mol}^\circ\text{C}}$$

Measuring and Expressing Heat Changes:

6. A student mixed 75.0 mL of water containing 0.75 mol HCl at 25°C with 75.0 mL of water containing 0.75 mol of NaOH at 25°C in a foam cup calorimeter. The temperature of the resulting solution increased to 35°C. How much heat in kilojoules was released by this reaction ($C_{\text{water}} = 4.18 \text{ J/g}^\circ\text{C}$)?

$$q = mc\Delta T$$

$$m = 150.0 \text{ g}$$

$$c = 4.184 \text{ J/g}^\circ\text{C}$$

$$\Delta T = 10^\circ\text{C}$$

$$q = (150.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(10^\circ\text{C})$$

$$\boxed{q = 6300 \text{ J}}$$

7. Calculate the amount of heat evolved when 150 g of water are reacted with excess calcium according to the following thermochemical equation:



$$\frac{150 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}} \times \frac{-65.2 \text{ kJ}}{1} = \boxed{-270 \text{ kJ}; \text{ exothermic}}$$

Thermochemistry Practice Problems 2 Answers

Kaplan Test Prep



Thermochemistry Practice Problems 2 Answers:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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