

# Thermochemistry Review

1. A reaction has  $\Delta H_{\text{rxn}}$  of  $+225 \text{ kJ/mol}$  and an activation energy of  $145 \text{ kJ/mol}$ . What is the activation energy for the reverse reaction?
2. Determine the enthalpies (including the enthalpy) of the formation of  $\text{H}_2\text{O}$  from its elements. Discuss the energy in and out. Draw the enthalpy diagram.
3. Determine the enthalpies (including enthalpy) of the decomposition of liquid water to hydrogen and oxygen. How much energy is required for making water?
4. Determine the molar enthalpy of combustion of ethanol  $\text{C}_2\text{H}_5\text{OH}$ . Compare it with the molar enthalpy of combustion of gasoline  $\text{C}_8\text{H}_{18}$  (octane).
5. What mass of pentane,  $\text{C}_5\text{H}_{12}$ , must be burned to release  $1.00 \text{ MJ}$  of energy?
6. What is the specific heat of an unknown metal if a metal mass of  $100.0 \text{ g}$  at  $100.0^\circ\text{C}$  is placed in  $100.0 \text{ g}$  of water at  $20.0^\circ\text{C}$ . What is the specific heat capacity of the metal?
7. What  $\Delta H_{\text{rxn}}$  of pentane is burned in burner at a rapid rate (weighting)  $20.0 \text{ g}$  of water inside a  $1.0 \text{ L}$  of water from  $20.0^\circ\text{C}$  to  $100.0^\circ\text{C}$  (boiling) in 10 minutes. What is the molar enthalpy of combustion of pentane? What is the enthalpy of the water?
8. The enthalpy of combustion of ethanol is  $-1366 \text{ kJ/mol}$ . Write the equation below:



Using Hess' law, prove that the formation enthalpy of ethanol is  $-277 \text{ kJ/mol}$  and water is  $-285 \text{ kJ/mol}$  for the reaction below for an ethyl alcohol:  $\text{C}_2\text{H}_5\text{OH}(l)$

# Thermochemistry Review Key

**M Mark**



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