

## CHAPTER 7

### THERMOCHEMISTRY SHORT QUESTION AND ANSWERS

Q.1 Define the following terms and give three examples of each.

Ans.

(i) System:

The substance which is under experiment or under observation is called as system.

Examples:

(i)  $\text{Pb}(\text{NO}_3)_2$  in decomposition of  $\text{Pb}(\text{NO}_3)_2$ .

(ii) Zn and  $\text{CuSO}_4$  solution, the reaction mixture in the vessel.

(iii)  $\text{CaCO}_3$  in thermal decomposition of  $\text{CaCO}_3$

(ii) Surroundings:

Everything around the system which is not a part of system is called surroundings.

For example

During the reaction between Zn and  $\text{CuSO}_4$  solution vessel and air etc are surroundings.

(iii) State function:

A macroscopic property of a system which has some definite value for initial and final state and independent of the path followed e.g.

(i) Pressure (ii) Temperature (iii) Internal energy.

Note: Heat is not a state function.

Q.2 Describe the units of energy.

Ans.

Mostly Joule and calorie are used for the measurement of energy.

Calorie:

The amount of energy required to raise the temperature of one gram of water from  $14.5^\circ\text{C}$  to  $15.5^\circ\text{C}$  is called one calorie.

Joule:

# Thermochemistry Short Answers

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adiabatic conditions for reversible process carnot s cycle its efficiency thermodynamic scale of temperature 5L C Thermochemistry Heat of reaction standard states relation between heat of reaction at constant volume at constant pressure Hess s law of constant heat of summation its applications bond dissociation energy its calculations from thermochemical data variation of heat of reaction with temperature Kirchoff s equation 4L Unit II Thermodynamics II A Second law of thermodynamics Need for second law of thermodynamics statements of second law of thermodynamics concept of entropy entropy as a state function of V T P T entropy change in phase change for ideal gas entropy as criteria of spontaneity equilibrium 4L B Free energy functions Helmholtz free energy A Gibb s free energy G their properties standard free energies effect of temperature on free energy Gibb s Helmholtz equation its applications A G as criteria for thermodynamic equilibrium 4L C System of variable composition Partial molar quantities chemical potential Van t Hoff s reaction isotherm relation between standard free energy change equilibrium constant thermodynamic derivation of law of mass action effect of temperature on equilibrium constant reaction isochore 4L Unit III Phase Equilibria A Phase rule Statement of phase rule definition of phase component and degree of freedom derivation of phase rule Clapeyron equation its application in deciding slopes of line for two phase equilibria applications of phase rule to two phase equilibria of i water system ii sulphur system iii Pb Ag system 6L B Liquid Liquid mixtures Ideal liquid mixtures Raoult s law of ideal solutions Henry s law non ideal systems azeotropes HCl H<sub>2</sub>O ethanol water system Partial miscible liquids Phenol water system trimethylamine water nicotine water system lower upper consolute temperature effect of impurity Immiscible liquids Steam distillation Nernst distribution law Limitations deviations applications 6L Unit IV Solid State Laws of crystallography i Law of constancy of interfacial angles ii Law of rationality of indices iii Law of symmetry symmetry of elements in crystals Unit cell space lattice orientation of lattice plane Miller indices Bravais lattices crystal systems X ray diffraction by crystal derivation of Braggs equation

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## **Table of Contents Thermochemistry Short Answers**

1. Understanding the eBook Thermochemistry Short Answers
  - The Rise of Digital Reading Thermochemistry Short Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Short Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Thermochemistry Short Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Short Answers
  - Personalized Recommendations

- Thermochemistry Short Answers User Reviews and Ratings
- Thermochemistry Short Answers and Bestseller Lists
- 5. Accessing Thermochemistry Short Answers Free and Paid eBooks
  - Thermochemistry Short Answers Public Domain eBooks
  - Thermochemistry Short Answers eBook Subscription Services
  - Thermochemistry Short Answers Budget-Friendly Options
- 6. Navigating Thermochemistry Short Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Thermochemistry Short Answers Compatibility with Devices
  - Thermochemistry Short Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Thermochemistry Short Answers
  - Highlighting and Note-Taking Thermochemistry Short Answers
  - Interactive Elements Thermochemistry Short Answers
- 8. Staying Engaged with Thermochemistry Short Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Thermochemistry Short Answers
- 9. Balancing eBooks and Physical Books Thermochemistry Short Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Thermochemistry Short Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Short Answers
  - Setting Reading Goals Thermochemistry Short Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Short Answers
  - Fact-Checking eBook Content of Thermochemistry Short Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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