

Thermochemistry

Unit 12 Page 1 (Use with Notes Page 1)

Directions: Determine whether the following scenarios best describe an endothermic, or exothermic process.

1. On a sunny winter day, the snow on a rooftop begins to melt. As the melted water drips from the roof, it refreezes into icicles. Describe the direction of the heat flow as the water freezes. Is this process endothermic or exothermic?
2. A container of melted wax stands at room temperature. What is the direction of heat flow as the liquid wax solidifies? Is the process endothermic or exothermic?
3. When barium hydroxide octahydrate, $\text{Ba}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$ is mixed in a beaker with ammonium thiocyanate, NH_4SCN , a reaction occurs. The beaker becomes very cold. Is the reaction endothermic or exothermic?
4. Two substances in a glass beaker chemically react, and the beaker becomes too hot to touch. Is the reaction endothermic or exothermic?

Directions: Make the following conversions between the different units for heat energy.

- | | |
|---|-------------------------------------|
| 5. 8.50×10^2 cal to kilocalories | 7. 1.8kJ to joules |
| 6. 444 cal to joules | 8. 4.5×10^3 kJ to calories |

Directions: Use your understanding of the specific heat formula to answer the following questions.

9. The temperature of a 95.4g piece of copper increases from 25.0°C to 48.0°C when the copper absorbs 849J of heat. What is the specific heat of copper?
10. When 435J of heat is added to 3.4g of olive oil at 21°C , the temperature increases to 85°C . What is the specific heat of the olive oil?
11. How much heat is required to raise the temperature of 250.0g of mercury 52°C if the specific heat of mercury is $0.14\text{J}/(\text{g} \cdot ^\circ\text{C})$?
12. How much heat is required to raise the temperature of 400.0g of silver 45°C if the specific heat of silver is $0.24\text{J}/(\text{g} \cdot ^\circ\text{C})$?

Thermochemistry

Thermochemistry Worksheet Answers Unit 12

PT Brinkman



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Table of Contents Thermochemistry Worksheet Answers Unit 12

1. Understanding the eBook Thermochemistry Worksheet Answers Unit 12
 - The Rise of Digital Reading Thermochemistry Worksheet Answers Unit 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Worksheet Answers Unit 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Worksheet Answers Unit 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Worksheet Answers Unit 12
 - Personalized Recommendations

- Thermochemistry Worksheet Answers Unit 12 User Reviews and Ratings
- Thermochemistry Worksheet Answers Unit 12 and Bestseller Lists
- 5. Accessing Thermochemistry Worksheet Answers Unit 12 Free and Paid eBooks
 - Thermochemistry Worksheet Answers Unit 12 Public Domain eBooks
 - Thermochemistry Worksheet Answers Unit 12 eBook Subscription Services
 - Thermochemistry Worksheet Answers Unit 12 Budget-Friendly Options
- 6. Navigating Thermochemistry Worksheet Answers Unit 12 eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Worksheet Answers Unit 12 Compatibility with Devices
 - Thermochemistry Worksheet Answers Unit 12 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Worksheet Answers Unit 12
 - Highlighting and Note-Taking Thermochemistry Worksheet Answers Unit 12
 - Interactive Elements Thermochemistry Worksheet Answers Unit 12
- 8. Staying Engaged with Thermochemistry Worksheet Answers Unit 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Worksheet Answers Unit 12
- 9. Balancing eBooks and Physical Books Thermochemistry Worksheet Answers Unit 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Worksheet Answers Unit 12
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Worksheet Answers Unit 12
 - Setting Reading Goals Thermochemistry Worksheet Answers Unit 12
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Worksheet Answers Unit 12
 - Fact-Checking eBook Content of Thermochemistry Worksheet Answers Unit 12

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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