

Thermodynamics Formulas

Name	Formula	
1 st Law of Thermodynamics	$\Delta E_{\text{universe}} = \Delta E_{\text{system}} + \Delta E_{\text{surroundings}}$	Energy cannot be created nor destroyed – only change forms.
2 nd Law of Thermodynamics	$\Delta S_{\text{universe}} = \Delta S_{\text{sys}} + \Delta S_{\text{sur}} \geq 0$	Entropy of the universe is always increasing.
3 rd Law of Thermodynamics	$S = K_B \ln W = K_B \ln(1) = 0$	S of a pure crystal at 0K, is 0 J/mol·K
Enthalpy	$\Delta H = q$	Enthalpy, ΔH , is the measure of heat energy in a system at constant pressure.
Standard Enthalpy	$\Delta H^{\circ} = \sum \Delta H^{\circ}_{f, \text{products}} + \sum \Delta H^{\circ}_{f, \text{reactants}}$	Enthalpy change when one mole of a substance is made from its elements in standard conditions. (Enthalpy Change of Formation)
Entropy	$\Delta S = \frac{q_{\text{reversible}}}{T}$	Entropy, ΔS , is the measure of disorder in a system. A higher entropy = A greater amount of disorder. Solids < Liquids < Gases.
	$\Delta S_{\text{surroundings}} = - \frac{\Delta H_{\text{system}}}{T}$	
Standard Entropy	$\Delta S^{\circ} = \sum \Delta S^{\circ}_{\text{products}} + \sum \Delta S^{\circ}_{\text{reactants}}$	(Standard Molar Entropies)
Gibb's Free Energy	$\Delta G = \sum \Delta G_{\text{products}} + \sum \Delta G_{\text{reactants}}$	Used to to predict whether a process will occur spontaneously at a constant temperature and pressure.
	$\Delta G = \Delta H - T \Delta S$	
	$\Delta G = \Delta G^{\circ} + RT \cdot \ln(Q)$	Q = Reaction Quotient R = 8.3145 J/mol·K
Standard Free Energy	$\Delta G^{\circ} = \sum \Delta G^{\circ}_{f, \text{products}} + \sum \Delta G^{\circ}_{f, \text{reactants}}$	
Specific Heat Capacity, c	$q = mc \Delta T$	The amount of heat energy needed to raise the temperature of 1 gram of a substance by 1°C. Used in "coffee-cup" calorimetry. m = mass in grams. (Specific Heat Capacities)
Heat Capacity, C	$q = C \Delta T$	The amount of energy required to raise the temperature of a substance by 1°C. Used in "bomb" calorimetry.

Thermodynamics Equations Cheat Sheet

Xiaolong Qi

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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