

3.3 Thermodynamics: Study Guide

Use this [outline](#) to review the key vocabulary for this unit.

1. Match the vocabulary words on the left with the appropriate definitions and units on the right.

a. Thermal Energy	<u>A</u> Kinetic energy in transit from one object to another due to temperature difference
b. Temperature	<u>B</u> Degree Celsius
	<u>C</u> Joules
	<u>D</u> The average kinetic energy of particles in an object

2. In the table below, write in the appropriate formulas.

Energy Transfer	Rate of Energy Transfer	Thermal Conductivity	Coefficient of Heat Conductivity
$Q = mc \cdot \Delta T$	$P = \frac{c \Delta T}{L}$ or $P = \frac{Q}{t}$ <u>Thermal conductance</u> $P = \sigma A c (\Delta T)^2$	$K = \frac{PL}{A \Delta T}$	$U = \frac{P}{A \cdot \Delta T}$

3. For each example, write the correct number for the law of thermodynamics it describes (0, 1, 2).
- 2 The law of entropy
 - 1 Thermal energy can only be increased in a system by adding heat from an external source or performing work within the system
 - 0 If two systems are separately found to be in thermal equilibrium with a third, then the two systems are in thermal equilibrium with each other
 - 2 Thermal energy flows only from warmer to cooler
 - 1 The law of conservation of energy applied to thermal systems
4. How much thermal energy would it take to raise the temperature of 1.1 (1 kg) of water by 10°C? The specific heat capacity of water is $4184 \frac{J}{kg \cdot ^\circ C}$

$$Q = mc \Delta T$$

$$= 1.1 kg \cdot 4184 \frac{J}{kg \cdot ^\circ C} \cdot 10^\circ C$$

$$Q = 41,840 J$$

$$Q = \underline{41,840 J}$$

5. A materials scientist is trying to determine the specific heat capacity of an experimental material. She applies 30 kJ of thermal energy to a 5 kg sample of the material and records a 24°C temperature increase in the sample. What is the specific heat capacity of the material?

$$Q = mc \Delta T$$

$$c = \frac{Q}{m \Delta T}$$

$$c = 2500 \frac{J}{kg \cdot ^\circ C}$$

$$c = \underline{2500 \frac{J}{kg \cdot ^\circ C}}$$

$$= \frac{30000 J}{5 kg \cdot 24^\circ C}$$

Thermodynamics Study Guide Answers

**Bennett Logan Huffman, Alton Lee
Scott**



Thermodynamics Study Guide Answers:

Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications *Organic Chemistry Study Guide with Solutions Manual* Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer **Study Guide/Solutions Manual for Organic Chemistry** K. Peter C. Vollhardt, Neil E. Schore, 2018-01-12 Updated for the Eighth Edition of Vollhardt Schore Organic Chemistry and written by the book's coauthor Neil Schore this invaluable manual includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer **Thermodynamics Problem Solving in Physical Chemistry** Kathleen E. Murphy, 2020-03-23 Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation create appropriate solutions and gain confidence through practice solving physical chemistry problems The workbook includes six major sections with 20-30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area Key Features Includes a visual map that shows how all the equations used in thermodynamics are connected and how they are derived from the three major energy laws Acts as a guide in deriving the correct solution to a problem Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry Can be used as a stand alone product for review of Thermodynamics questions for major tests **Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry** Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface *Thermodynamics* Ibrahim Dincer, 2020-08-18 Presents a unique stepwise exergy based approach to thermodynamic concepts systems and applications Thermodynamics A Smart Approach redefines this crucial branch of engineering as the science of energy and exergy rather than the science of energy and entropy to provide an innovative step by step approach for teaching understanding and practicing thermodynamics in a clearer and easier way Focusing primarily on the concepts and balance equations this innovative textbook covers exergy

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Biomass as a Sustainable Energy Source for the Future Wiebren de Jong,J. Ruud van Ommen,2014-10-03 Focusing on the conversion of biomass into gas or liquid fuels the book covers physical pre treatment technologies thermal chemical and biochemical conversion technologies Details the latest biomass characterization techniques Explains the biochemical and thermochemical conversion processes Discusses the development of integrated biorefineries which are similar to petroleum refineries in concept covering such topics as reactor configurations and downstream processing Describes how to mitigate the environmental risks when using biomass as fuel Includes many problems small projects sample calculations and industrial application examples

Thermodynamics Stephen R. Turns,Laura L. Pauley,2020-02-27 Fully revised to match the more traditional sequence of course materials this full color second edition presents the basic principles and methods of thermodynamics using a clear and engaging style and a wealth of end of chapter problems It includes five new chapters on topics such as mixtures psychrometry chemical equilibrium and combustion and discussion of the Second Law of Thermodynamics has been expanded and divided into two chapters allowing instructors to introduce the topic using either the cycle analysis in Chapter 6 or the definition of entropy in Chapter 7 Online ancillaries including new LMS testbanks a password protected solutions manual prepared PowerPoint lecture slides instructional videos and figures in electronic format are available at www.cambridge.org/thermo

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