

Thermodynamics IN-CLASS Practice Assessment:

Skill 1 - Can define and identify direction of heat flow for exothermic and endothermic situations

- Using the image to the right, draw arrows and describe as many endothermic and exothermic processes as possible.



Skill 2 - Use the equation $\Delta E = q + w$ correctly when given these terms in a real world situation

- What is the ΔE if a substance releases 100J of energy and has 50J of work done on it?
- How many calories of energy correspond to 28.4 J?

Skill 3 - Use $q = cm\Delta T$ to calculate the missing quantity, given the others

- Using p328, identify the substance being described if 42.6 J is required to heat 2.0 g of this substance from 50°C to 60°C.
- Calculate the number of Calories needed to raise the temperature of 275.3 g of water from 17.3°C to 63.7°C?

Skill 4 - Calculate the energy of a reaction using Hess's Law and identify if it is exothermic or endothermic

- Given the following data:

$\text{CO}_2 \rightarrow \text{C} + \text{O}_2$	$\Delta H = 393.5 \text{ kJ}$
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What is the ΔH for:



$\Delta H =$

Exothermic or

Endothermic?

- Given the following data, calculate the ΔH for the formation of ethane (C_2H_6)

$\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$	$\Delta H = -393.5 \text{ kJ}$
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Thermodynamics Test Review Key

Joacim Rocklöv

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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