

Percent Yield Answer Key

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

Actual yield is the experimental value.

Theoretical yield is the calculated value.

6. Methylcyclopentane reacts with bromine to produce bromocyclopentane. 1.50 g of methylcyclopentane is used. Calculate the percent yield of bromocyclopentane.



$$1.50 \text{ g Methylcyclopentane} \times \frac{1 \text{ mol Methylcyclopentane}}{98.15 \text{ g Methylcyclopentane}} \times \frac{1 \text{ mol Br}_2}{1 \text{ mol Methylcyclopentane}} \times \frac{1 \text{ mol bromocyclopentane}}{1 \text{ mol Br}_2} \times \frac{134.17 \text{ g bromocyclopentane}}{1 \text{ mol bromocyclopentane}} = 2.00 \text{ g bromocyclopentane}$$

$$\frac{1.50 \text{ g bromocyclopentane}}{2.00 \text{ g bromocyclopentane}} \times 100\% = 75.0\% \text{ yield}$$

7. Acetylene is produced by reacting calcium carbide with water. Acetylene reacts with oxygen to produce CO_2 and water. 10.0 g of acetylene reacts with oxygen to form carbon dioxide. Calculate the percent yield of this reaction.



$$10.0 \text{ g C}_2\text{H}_2 \times \frac{1 \text{ mol C}_2\text{H}_2}{26.04 \text{ g C}_2\text{H}_2} \times \frac{4 \text{ mol CO}_2}{2 \text{ mol C}_2\text{H}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 33.3 \text{ g CO}_2$$

$$\frac{25.0 \text{ g CO}_2}{33.3 \text{ g CO}_2} \times 100\% = 75.1\% \text{ yield}$$

8. Ethylacetate reacts with sodium hydroxide to produce sodium acetate and ethanol. 10.0 g of sodium acetate reacts to form 4.0 g of ethanol. Calculate the percent yield of this reaction.



$$10.0 \text{ g C}_2\text{H}_5\text{CO}_2\text{Na} \times \frac{1 \text{ mol C}_2\text{H}_5\text{CO}_2\text{Na}}{98.07 \text{ g C}_2\text{H}_5\text{CO}_2\text{Na}} \times \frac{1 \text{ mol C}_2\text{H}_5\text{OH}}{1 \text{ mol C}_2\text{H}_5\text{CO}_2\text{Na}} \times \frac{46.07 \text{ g C}_2\text{H}_5\text{OH}}{1 \text{ mol C}_2\text{H}_5\text{OH}} = 4.70 \text{ g C}_2\text{H}_5\text{OH}$$

$$\frac{4.0 \text{ g C}_2\text{H}_5\text{OH}}{4.70 \text{ g C}_2\text{H}_5\text{OH}} \times 100\% = 85.1\% \text{ yield}$$

9. Ethanol can be produced through the reaction of CO and H_2 using a catalyst. 10.0 g of CO reacts to produce 8.0 g of ethanol. Calculate the percent yield of $\text{C}_2\text{H}_5\text{OH}$.



$$10.0 \text{ g CO} \times \frac{1 \text{ mol CO}}{28.01 \text{ g CO}} \times \frac{1 \text{ mol C}_2\text{H}_5\text{OH}}{2 \text{ mol CO}} \times \frac{46.07 \text{ g C}_2\text{H}_5\text{OH}}{1 \text{ mol C}_2\text{H}_5\text{OH}} = 8.20 \text{ g C}_2\text{H}_5\text{OH}$$

$$\frac{8.0 \text{ g C}_2\text{H}_5\text{OH}}{8.20 \text{ g C}_2\text{H}_5\text{OH}} \times 100\% = 97.6\% \text{ yield}$$

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