

Check for Understandings: Atomic Theory - Structure (HS-PS1-1)

Name _____

Atomic Structure

1) True or False? If False, correct the statement.

_____ a. J.J. Thompson assumed that the atom must contain positively charged particles, in addition to negatively charged particles because isolated atoms have no overall charge.

_____ b. In the plum pudding model, the atom was envisioned as a sphere of negative charge in which positively charged electrons were randomly distributed.

_____ c. Rutherford's bombardment experiments with gold foil suggested that alpha particles were deflected by approaching a negatively charged atomic nucleus.

_____ d. The proton and neutron have similar relative masses but different charges.

_____ e. Atoms that have the same number of neutrons, but different numbers of protons are called isotopes.

2) Strontium-90 occurs in fallout from nuclear testing. It can accumulate in bone marrow and may cause leukemia and bone cancer. Give the number of protons, neutrons, and electrons in atom symbolized by $^{90}_{38}\text{Sr}$.

3) Copy and complete the following table.

Name	Symbol	Atomic Number	Mass Number	Neutrons	Electrons
Sodium		11	23		
	$^{15}_7\text{N}$				
	$^{238}_{92}\text{U}$				
Lithium			7		3
Boron		5	11		

Theory Of Atomic Structure Answer Key

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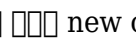





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