

CHEMISTRY Terminology

The ABCs of chemistry

A

A: Symbol for atomic mass number
absolute zero: Lowest temperature; 0 Kelvin
Ae: Symbol for the element actinium, $Z = 89$
accuracy: The closeness of measurement to true value
acetate anion: $C_2H_3O_2^-$, charge -1
acetic acid: $C_2H_4O_2$
acetone: Dimethyl ketone; organic solvent
acid-base indicator: A chemical that changes color if the pH changes
acid-base titration: A method for determining acid or base concentrations
acid definition: See Arrhenius acid, Brønsted-Lowry acid, Lewis acid
acid ionization constant (K_a): Equilibrium constant for acid dissociation
acidic solution: pH below 7
actinide: Element with $Z = 90$ or above; radioactive
activation energy (E_a): A process's energy barrier



adhesion: Attraction of unlike molecules or materials
Ag: Symbol for the element silver, $Z = 47$
Al: Symbol for the element aluminum, $Z = 13$
alcohol: Organic compound with -OH group; ROH
aldehyde: Carbonyl group bonded to hydrogen and an organic group
alkali metal family: Lithium, sodium, potassium, rubidium, cesium, and francium; column #1
alkaline earth metal family: Beryllium, magnesium, calcium, strontium, barium, and radium; column #2
alkane: Hydrocarbon; all C-C single bonds
alkene: Hydrocarbon; 1 or more C=C double bonds
alkyne: Hydrocarbon; 1 or more C≡C triple bonds
allotropes: Two or more forms of an element
alloy: Solution of 2 or more metals
alpha (α): Greek letter that denotes radioactive particle and various scientific variables
alpha particle: Helium nucleus; charge +2
Au: Symbol for the element auriferous, $Z = 79$
amalgam: Alloy of mercury with other metal
amine: Organic base; RRR'N; modified ammonia
amino acid: Compound with organic acid and organic base property; forms proteins and peptides
ammonia: NH_3 , base
ammonium cation: NH_4^+ , charge +1
ammonium hydroxide: NH_4OH
amphoteric oxide: Exhibits both acid and basic properties
amu: Atomic mass unit; $1/12$ mass of C-12
anion: Ion with a negative charge
anode: Electrode that supports oxidation
antibonding MO: MO is less stable than separate AOs
AO: Abbreviation for atomic orbital

aq: Aqueous; "dissolved in water"
Ar: Symbol for the element argon, $Z = 18$
aromatic: Organic compound with a benzene ring
Arrhenius acid: Produces hydronium ion in water solution
Arrhenius base: Yields hydroxide ion in water solution
Arrhenius equation for rate constant (k):

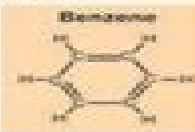
$$k = Ae^{-E_a/RT}$$
 E_a = activation energy
arsenide anion: As^{3-} , charge -3
As: Symbol for the element arsenic, $Z = 33$
At: Symbol for the element astatine, $Z = 85$
atm: Symbol for pressure in "atmospheres"
atom: Fundamental unit of all matter
atomic mass number (A): Total number of protons and neutrons
atomic number (Z): Number of protons in the nucleus
atomic orbital: Wave motion of electrons in atoms
atomic radius: Empirical measure of atom size
atomic weight: Weighted average of natural isotopes of an element
Au: Symbol for the element gold, $Z = 79$
Aufbau principle: Guides the filling of electronic subshells of the elements
Avogadro's law: Volume is proportional to gas moles at fixed pressure and temperature
Avogadro's number: $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

Bk: Symbol for the element berkelium, $Z = 97$
bleach: Chemical that is a strong oxidizing agent
boiling: Liquid $\rightarrow$ gas at the boiling point
boiling point elevation: A solution has a higher boiling point than a pure solvent
boiling point (T_b): Liquid-gas equilibrium, $P = 1 \text{ atm}$
boiling point of water: 100°C
bond energy: Energy held by a chemical bond
bond length: Distance between 2 atoms in a bond
bond order for Lewis model: # of bonds divided by # of bonded atoms
bond order for MO treatment: # of filled bonding MOs minus # of filled antibonding MOs
bonding electrons: Form chemical bonds in a compound
bonding MO: MO is more stable than AOs
Boyle's law: PV constant for gas, fixed T and n
Br: Symbol for the element bromine, $Z = 35$
bromide anion: Br^- , charge -1
Brønsted-Lowry acid: Proton donor
Brønsted-Lowry base: Proton acceptor
buffer: Composed of weak acid and weak base; serves to keep pH constant



B

B: Symbol for the element boron, $Z = 5$
Ba: Symbol for the element barium, $Z = 56$
balanced equation: Each side of the equation has the same number of atoms of each element and charge
balanced redox: Electron loss = electron gain
barium cation: Ba^{2+} , charge +2
base definition: See Arrhenius base, Brønsted-Lowry base, Lewis base
base ionization constant (K_b): Equilibrium constant for base dissociation
basic solution: pH over 7
battery: Source of electrical power; galvanic cell
Be: Symbol for the element beryllium, $Z = 4$
benzene: C_6H_6 ; organic solvent; aromatic ring of 6 carbon atoms
beryllium cation: Be^{2+} , charge +2
beta (β): Greek letter that denotes radioactive particle and a number of scientific variables
beta particle: Energetic electron from the nucleus; charge -1
Bh: Symbol for the element bohrium, $Z = 107$
Bi: Symbol for the element bismuth, $Z = 83$
bicarbonate anion: HCO_3^- , charge -1
bimolecular oxygen: O_2
binary compound: Formed from 2 elements
bisulfate anion: HSO_4^- , charge -1



b: Symbol for centi- , the SI prefix for $1/100$ or 10^{-2}
c: Symbol for the speed of light
°C: Celsius temperature scale
C: Symbol for the element carbon, $Z = 6$
 $C_2H_4O_2$: Acetic acid
Ca: Symbol for the element calcium, $Z = 20$
calcium cation: Ca^{2+} , charge +2
calcium fluoride: CaF_2
calcium hydroxide: $Ca(OH)_2$
calorie: English unit of heat energy
calorimetry: Study of the release or absorption of heat
carbide anion: C^{4-} , charge -4
carbohydrate: Organic compound; carbon bonded to several -OH groups (sugar and starch)
carbon dioxide: CO_2
carbon monoxide: CO
carbonate anion: CO_3^{2-} , charge -2
carbonic acid: H_2CO_3
carbonyl group: CO , ligand complex
carboxyl group: $-CO$ bonded to 2 groups
carboxylic acid: Organic acid; $R-COOH$
catalyst: Accelerates a reaction but is not consumed in the reaction; lowers the activation energy
cathode: Electrode that supports reduction
cation: Ion with a positive charge
Cd: Symbol for the element cadmium, $Z = 48$



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R Bogdan



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Polymers have never been fully integrated into the disciplines of science and engineering. This book is designed to teach thermal analysis and the understanding of all materials flexible macromolecules as well as those of the small molecules and rigid macromolecules. The macroscopic tool of inquiry is thermal analysis and the results are linked to microscopic molecular structure and motion. Measurements of heat and mass are the two roots of quantitative science. The macroscopic heat is connected to the microscopic atomic motion while the macroscopic mass is linked to the microscopic atomic structure. The macroscopic units of measurement of heat and mass are the joule and the gram chosen to be easily discernable by the human senses. The microscopic units of motion and structure are 12 10 the picosecond 10 seconds and the nanogram 10 meters chosen to fit the atomic scales. One notes a factor of 10 000 between the two atomic units when expressed in human units second and gram with one gram being equal to one cubic centimeter when considering water. Perhaps this is the reason for the much better understanding and greater interest in the structure of materials being closer to human experience when compared to molecular motion.

An Elementary manual of chemistry Francis Humphreys Storer, 1894 An Elementary Manual of Chemistry William Ripley Nichols, William B. Lindsay, Frank Humphreys Storer, 1894 Monthly Catalog of United States Government Publications, 1984 Proceedings of the 4th International Conference on Industrial Engineering Andrey A. Radionov, Oleg A. Kravchenko, Victor I. Guzeev, Yuriy V. Rozhdestvenskiy, 2018-12-07. This book highlights recent findings in industrial manufacturing and mechanical engineering and provides an overview of the state of the art in these fields mainly in Russia and Eastern Europe. A broad range of topics and issues in modern engineering are discussed including the dynamics of machines and working processes friction wear and lubrication in machines surface transport and technological machines manufacturing engineering of industrial facilities materials engineering metallurgy control systems and their industrial applications industrial mechatronics automation and robotics. The book gathers selected papers presented at the 4th International Conference on Industrial Engineering ICIE held in Moscow Russia in May 2018. The authors are experts in various fields of engineering and all papers have been carefully reviewed. Given its scope the book will be of interest to a wide readership including mechanical and production engineers lecturers in engineering disciplines and engineering graduates.

Energy Research Abstracts, 1989 Trade and Industrial Education, 1972 Trade and Industrial Education; Instructional Materials Ohio State University. Center for Vocational and Technical Education, 1972

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Table of Contents Thermal Chemistry Study Guide

1. Understanding the eBook Thermal Chemistry Study Guide
 - The Rise of Digital Reading Thermal Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermal Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermal Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermal Chemistry Study Guide
 - Personalized Recommendations
 - Thermal Chemistry Study Guide User Reviews and Ratings
 - Thermal Chemistry Study Guide and Bestseller Lists
5. Accessing Thermal Chemistry Study Guide Free and Paid eBooks

- Thermal Chemistry Study Guide Public Domain eBooks
- Thermal Chemistry Study Guide eBook Subscription Services
- Thermal Chemistry Study Guide Budget-Friendly Options
- 6. Navigating Thermal Chemistry Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermal Chemistry Study Guide Compatibility with Devices
 - Thermal Chemistry Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermal Chemistry Study Guide
 - Highlighting and Note-Taking Thermal Chemistry Study Guide
 - Interactive Elements Thermal Chemistry Study Guide
- 8. Staying Engaged with Thermal Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermal Chemistry Study Guide
- 9. Balancing eBooks and Physical Books Thermal Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermal Chemistry Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermal Chemistry Study Guide
 - Setting Reading Goals Thermal Chemistry Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermal Chemistry Study Guide
 - Fact-Checking eBook Content of Thermal Chemistry Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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