

## CHAPTER

## 1. Exothermic and Endothermic reactions

For a given chemical change



Change in internal energy during reaction ( $\Delta U$ )

= Internal energy of products – Internal energy of reactants

i.e.,  $\Delta U = \Sigma U_P - \Sigma U_R$

if  $\Delta U = +ve$ ; the reaction is endothermic

if  $\Delta U = -ve$ ; the reaction is exothermic

The exothermic and endothermic nature are expressed as

Exothermic :  $C + O_2 \longrightarrow CO_2; \Delta U = -94.3 \text{ kcal}$

or  $C + O_2 \longrightarrow CO_2 + 94.3 \text{ kcal}$

Endothermic :  $C + 2S \longrightarrow CS_2; \Delta U = +22 \text{ kcal}$

or  $C + 2S \longrightarrow CS_2 - 22 \text{ kcal}$

## 2. Heat of reaction : It is defined as,

—the amount of heat absorbed or released (i.e., the change in internal energy or the change in heat enthalpy)

(a) Condition of constant  $P$  or  $V$  :

Heat changes at constant volume are expressed in  $\Delta U$ .

Heat changes at constant pressure are expressed in  $\Delta H$ .

Also, for a change  $\Delta H = \Delta U + P\Delta V$  —(1)

where,  $P$  is the pressure and  $\Delta V$  is change in volume.

Also,  $\Delta H$  and  $\Delta U$  are related together as

$$\Delta H = \Delta U + \Delta n \cdot RT$$
 —(2)

$\Delta n$  = Mole of products – Mole of reactants

$R$  = Molar gas constant

$T$  = Temperature in Kelvin.

# Thermochemistry Conceptual Questions

**Scott C. Dulebohn**



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