

CHEM 101 Worksheet, Chapter 6: Thermochemistry

Multiple choice questions 1-10 (page 1): circle the correct answer(A-E)

Show your work questions 11-13 (pages 2-3): give the complete answer in the space provided

R = 8.314 J/(mol.K) = 0.08206 L.atm/(mol.K) 1 cal = 4.184 J 0 °C = 273 K

1. Calculate the heat effect q (in kJ) when 112 g of iron (Fe(s)) is heated from 20 °C to 300 °C. The specific heat of Fe(s) = 0.48 J/g.K

- A. -34.9 kJ B. +34.9 kJ C. -15.0 kJ D. -270 J E. +15.0 kJ

2. The dissolution reaction for ammonium chloride: $\text{NH}_4\text{Cl(s)} \rightarrow \text{NH}_4^+(\text{aq}) + \text{Cl}^-(\text{aq})$ is endothermic. When ammonium chloride is dissolved in water in an insulated (adiabatic) system, the temperature will:

- A. increase B. decrease C. stay the same

3. Consider the reaction: $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO(s)} + \text{CO}_2(\text{g})$. $\Delta H_{\text{rxn}} = +178 \text{ kJ}$ What is the enthalpy change when 20.0 kg CaO is formed?

- A. $+6.36 \times 10^4 \text{ kJ}$ B. $-6.36 \times 10^4 \text{ kJ}$ C. $+6.36 \times 10^4 \text{ J}$ D. 63.6 kJ E. +178 kJ

4. Consider the reaction: $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$. $\Delta H_{\text{rxn}} = -114.6 \text{ kJ}$ As always, the ΔH is "for the reaction as written"! Calculate the enthalpy change ΔH , when 12.6 g NO_2 is produced.

- A. -31.4 kJ B. +31.4 kJ C. -15.7 kJ D. +15.7 kJ E. none of these

5. How much heat is associated with the reaction that produces 256.1 g of ammonia, NH_3 ?



- A. 92.45 kJ of heat is released B. 696.1 kJ of heat is released C. 184.9 kJ of heat is absorbed
D. 184.9 kJ of heat is released E. 92 kJ of heat is released

6. Which two of the following are not formation reactions?

1. $\frac{1}{2}\text{N}_2(\text{g}) + \frac{1}{2}\text{H}_2(\text{g}) \rightarrow \text{NH}_3(\text{g})$ 2. $\text{C(s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$ 3. $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
4. $2\text{Fe(s)} + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{Fe}_2\text{O}_3(\text{s})$ 5. $\frac{1}{2}\text{H}_2(\text{g}) + \frac{1}{2}\text{I}_2(\text{g}) \rightarrow \text{HI(g)}$

- A. 1 and 2 B. 3 and 4 C. 2 and 3 D. 4 and 5 E. 3 and 5

7. Consider the thermochemical reaction equation: $2\text{CO(g)} + \text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g})$ $\Delta H = -198 \text{ kJ}$

What is the enthalpy change for the reaction: $\text{CO}_2(\text{g}) \rightarrow \text{CO(g)} + \frac{1}{2}\text{O}_2(\text{g})$

- A. +198 kJ B. -99 kJ C. +99 kJ D. +396 kJ E. -198 kJ

8. Calculate the enthalpy change for the reaction: $2\text{KClO}_3 \rightarrow 2 \text{KCl} + \text{O}_2(\text{g})$

Standard enthalpy of formation values: $\Delta H_f^\circ(\text{KClO}_3) = -389 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{KCl}) = -436 \text{ kJ/mol}$

- A. -94 kJ B. -47 kJ C. +94 kJ D. +47 kJ
E. Cannot be calculated because $\Delta H_f^\circ(\text{O}_2(\text{g}))$ is not given

9. Given the following two thermochemical equations:



What is the enthalpy change for the reaction: $2\text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O(l)} + \text{O}_2(\text{g})$?

- A. -98 kJ B. +98 kJ C. +196 kJ D. -196 kJ E. -668 kJ

10. Which of the following statements is correct for a combustion reaction in an adiabatic "bomb" calorimeter?

- A. the pressure is constant B. the temperature is constant C. the volume is constant
D. the calorimeter constant's C units are J/g.K E. $\Delta T(\text{K}) = \Delta T(^{\circ}\text{C}) + 273.15$

Show your work for the following questions.

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Jin-Ying Zhang



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1. Understanding the eBook Thermochemistry Practice Multiple Choice Questions
 - The Rise of Digital Reading Thermochemistry Practice Multiple Choice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Practice Multiple Choice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Practice Multiple Choice Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Practice Multiple Choice Questions
 - Personalized Recommendations
 - Thermochemistry Practice Multiple Choice Questions User Reviews and Ratings

- Thermochemistry Practice Multiple Choice Questions and Bestseller Lists
- 5. Accessing Thermochemistry Practice Multiple Choice Questions Free and Paid eBooks
 - Thermochemistry Practice Multiple Choice Questions Public Domain eBooks
 - Thermochemistry Practice Multiple Choice Questions eBook Subscription Services
 - Thermochemistry Practice Multiple Choice Questions Budget-Friendly Options
- 6. Navigating Thermochemistry Practice Multiple Choice Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Practice Multiple Choice Questions Compatibility with Devices
 - Thermochemistry Practice Multiple Choice Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Practice Multiple Choice Questions
 - Highlighting and Note-Taking Thermochemistry Practice Multiple Choice Questions
 - Interactive Elements Thermochemistry Practice Multiple Choice Questions
- 8. Staying Engaged with Thermochemistry Practice Multiple Choice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Practice Multiple Choice Questions
- 9. Balancing eBooks and Physical Books Thermochemistry Practice Multiple Choice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Practice Multiple Choice Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Practice Multiple Choice Questions
 - Setting Reading Goals Thermochemistry Practice Multiple Choice Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Practice Multiple Choice Questions
 - Fact-Checking eBook Content of Thermochemistry Practice Multiple Choice Questions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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