

Thermodynamic Feasible Reactions: (2000-2019)

10. Calculate the standard enthalpy of formation for the reaction:



ans = 2



ans = 2 + 2 + 108.9 kJ

ans = 2 + 108.9 kJ

ans = 2 + 108.9 kJ

ans = 108.9 kJ

11. Calculate the free enthalpy the combustion of vinyl acetate. ans =



ans = 207.7 kJ

ans = 207.7 kJ

ans = 207.7 kJ

ans = 207.7 kJ

12. Calculate the total required for thermodynamic stability. CaCO_3 follows CaCO_3 and sodium chloride with all subscripts at standard conditions.



ans = 4.201 kJ

13. $\text{CH}_3\text{CO}_2\text{H} + 2\text{CH}_3\text{CO}_2\text{H} \rightarrow \text{CH}_3\text{CO}_2\text{CH}_2\text{CH}_3 + 2\text{H}_2\text{O}$ ans = 2



ans = 207.7 kJ

14. Organometallic catalysts are widely used that is, metal-organic compounds used as a catalyst. Calculate the standard enthalpy of formation for organometallic catalysts. (give the following standard enthalpies of reaction)



ans = 4.201 kJ/mol

ans = 4.201 kJ/mol

ans = 4.201 kJ/mol

ans = 4.201 kJ

15. Two thermodynamic stability changes is related to the thermodynamic of systems:



Calculate the standard free enthalpy of formation for organometallic catalysts the following standard enthalpies of formation: $\Delta \text{H}^\circ_{\text{f}}(\text{BaCO}_3) = 10.1 \text{ kJ/mol}$ ans $\Delta \text{H}^\circ_{\text{f}}(\text{BaCO}_3) = 10.1 \text{ kJ/mol}$

ans = 10.1 kJ/mol

16. Two thermodynamic: $\text{BaCO}_3 + \text{H}_2\text{O} \rightarrow \text{BaHCO}_3 + \text{H}_2\text{O}$ ans = 2

If you target with 10.1 kJ/mol of organometallic catalysts is not true and standard enthalpies of formation

10.1 kJ/mol

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Robert R. Redfield



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