

Name: _____

PROBLEM SET ____ THERMOCHEMISTRY

Answer the following exercises. For problem-solving, solutions should be written in a separate sheet.

1. Calculate the change in internal energy and determine whether the process is endothermic or exothermic for the following cases:

Situation	ΔE	Exothermic or Endothermic?
$q = 0.555 \text{ kJ}$ and $w = -950 \text{ J}$		
A system absorbs 75.0 kJ of heat to its surroundings while the surroundings do 58.0 kJ of work on the system.		
A balloon is cooled by removing 0.700 kJ of heat. It shrinks on cooling, and the atmosphere does 455 J of work on the balloon.		
A bar of gold is heated during which it absorbs 257 J of heat. Assume the volume of the gold bar remains constant.		

2. Indicate which of the following is independent of the path by which a change occurs:

Situation	Dependent or Independent?
the change in potential energy when a book is transferred from table to shelf	
the heat evolved when a cube of sugar is oxidized to $\text{CO}_2(\text{g})$ and $\text{H}_2\text{O}(\text{g})$	
the work accomplished in burning a gallon of gasoline	

3. Predict which of the following has the higher enthalpy in each case by putting > or < on the blanks provided.

	> or <	
1 mol $\text{I}_2(\text{s})$ at 20°C		1 mol $\text{I}_2(\text{g})$ at 20°C
2 mol of I atoms		1 mol of I_2
1 mol $\text{I}_2(\text{g})$ and 1 mol $\text{H}_2(\text{g})$ at 25°C		2 mol $\text{HI}(\text{g})$ at 25°C
1 mol $\text{H}_2(\text{g})$ at 100°C		1 mol $\text{H}_2(\text{g})$ at 300°C .

4. Consider the following reaction:



$\Delta H = -537 \text{ kJ}$



$\Delta H = -680 \text{ kJ}$



$\Delta H = +52.3 \text{ kJ}$

Which reaction(s) (1, 2, 3) is(are) exothermic?	
What is the ΔH for the reaction $\text{C}(\text{s}) + \text{H}_2(\text{g}) \rightarrow 1/2\text{C}_2\text{H}_4(\text{g})$?	
What is the ΔH for the reaction $4\text{CF}_4(\text{g}) \rightarrow 4\text{C}(\text{s}) + 8\text{F}_2(\text{g})$?	
What is the ΔH for the reaction $\text{C}_2\text{H}_4(\text{g}) + 6\text{F}_2(\text{g}) \rightarrow 2\text{CF}_4(\text{g}) + 4\text{HF}(\text{g})$?	
What is the ΔH for the reaction $3\text{CF}_4(\text{g}) + 6\text{HF}(\text{g}) \rightarrow 3/2\text{C}_2\text{H}_4(\text{g}) + 9\text{F}_2(\text{g})$?	
What is the ΔH for the reaction if 100.00 g of $\text{HF}(\text{g})$ decomposed completely to $\text{H}_2(\text{g})$ and $\text{F}_2(\text{g})$?	
What is the ΔH for the reaction if 100.00 g of $\text{C}(\text{s})$ reacted with excess $\text{F}_2(\text{g})$ to form $\text{CF}_4(\text{g})$?	
If the amount of heat resulting from the formation of $\text{C}_2\text{H}_4(\text{g})$ is +200.00 kJ, how many grams of $\text{H}_2(\text{g})$ reacted with excess $\text{C}(\text{s})$?	

Thermochemistry Problem Set One Specific Heat

Rosina Ehmann



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