

## Thermochemistry – Sol'ns & Explanations to Questions/Problems Con't

**Page 16 – Question:** Which is/are endothermic process(es)?

1. *Combustion* – a redox rxn that involves the rapid combination of a substance with  $O_2(g)$  to produce heat and usually light, i.e., ALWAYS exothermic!
2. *Fusion* – conversion of a solid to a liquid, at constant pressure and temperature; endothermic process
3. *Vaporisation* – conversion of a liquid to a gas; endothermic process
4. *Sublimation* – direct conversion of a solid to the gas phase; endothermic process
5. *Atomisation* – conversion of a substance into its constituent gaseous atoms; endothermic process

**Page 18 – Question:** The standard enthalpy of formation of sulfuric acid,  $H_2SO_4(l)$ , refers to the  $\Delta H^\circ$  of which of the following reactions?

- A)  $H_2SO_4(l) \rightarrow H_2(g) + S(s) + 2O_2(g)$  ✗ - compd on wrong side of equation  
 B)  $H_2(g) + S(g) + 2O_2(g) \rightarrow H_2SO_4(l)$  ✗ - sulfur not in its standard state  
 C)  $H_2SO_4(l) \rightarrow 2H(g) + S(s) + 4O(g)$  ✗ - compd on wrong side of equation  
 D)  $H_2(g) + S(s) + 2O_2(g) \rightarrow H_2SO_4(l)$  ✓ **More Correct:**  $H_2(g) + \frac{1}{8}S_8(r) + 2O_2(g) \rightarrow H_2SO_4(l)$   
 E)  $2H(g) + S(g) + 4O(g) \rightarrow H_2SO_4(l)$  ✗ - hydrogen & sulfur are not in their std. state

**Standard Enthalpy of Formation:** the change in enthalpy when one mole of a substance in the standard state (1 bar and 298.15 K) is formed from its pure elements under the same conditions.

**Page 20 – Question:** Calculate the heat required to decompose limestone:  $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$

Given:  $\Delta H^\circ_f[CaCO_3(s)] = -1206.9 \text{ kJ/mol}$ ;  $\Delta H^\circ_f[CaO(s)] = -635.1 \text{ kJ/mol}$ ;  $\Delta H^\circ_f[CO_2(g)] = -393.5 \text{ kJ/mol}$

1) **The Long Way (Hess's Law):**



Want:  $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$  Therefore, flip equation (1) and add (2) and (3):



2) **The Short Way:**

$$\begin{aligned}\Delta H^\circ_{rxn} &= \Sigma n\Delta H^\circ_f(\text{products}) - \Sigma n\Delta H^\circ_f(\text{reactants}) \\ \Delta H^\circ_{rxn} &= \Delta H^\circ_f[CaO(s)] + \Delta H^\circ_f[CO_2(g)] - \Delta H^\circ_f[CaCO_3(s)] \\ \Delta H^\circ_{rxn} &= (-635.1 \text{ kJ/mol}) + (-393.5 \text{ kJ/mol}) - (-1206.9 \text{ kJ/mol}) \\ \Delta H^\circ_{rxn} &= +178.3 \text{ kJ/mol}\end{aligned}$$

**Page 22 – Problem 1:** Calculate  $\Delta H^\circ_{rxn}$  for the reaction:  $CH_4(g) + 2O_2(g) \rightarrow CO_2(g) + 2H_2O(l)$

$$\begin{aligned}\Delta H^\circ_{rxn} &= \Sigma n\Delta H^\circ_f(\text{products}) - \Sigma n\Delta H^\circ_f(\text{reactants}) \\ \Delta H^\circ_{rxn} &= \Delta H^\circ_f[CO_2(g)] + 2\Delta H^\circ_f[H_2O(l)] - \Delta H^\circ_f[CH_4(g)] - 2\Delta H^\circ_f[O_2(g)] \\ \Delta H^\circ_{rxn} &= (-393.5 \text{ kJ/mol}) + 2(-285.8 \text{ kJ/mol}) - (-74.87 \text{ kJ/mol}) - 2(0 \text{ kJ/mol}) \\ \Delta H^\circ_{rxn} &= -890 \text{ kJ/mol}\end{aligned}$$

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Problems in Chemical Thermodynamics with Solutions Maka Aleksishvili, Shota Sidamonidze, 2002 The methods of chemical thermodynamics are effectively used in many fields of science and technology Mastering these methods and their use in practice requires profound comprehension of the theoretical questions and acquisition of certain calculating skills This book is useful to undergraduate and graduate students in chemistry as well as chemical thermal and refrigerating technology it will also benefit specialists in all other fields who are interested in using these powerful methods in their practical activities

**Basic Thermochemistry in Materials Processing** Gabriel Plascencia, David Jaramillo, 2017-02-28 This book provides the reader with some thermochemistry notes The intention is to provide a simple easy to understand text which serves as a complimentary material to more complex books It also provide students and those beginning in the field with several application examples used in different areas of materials processing The book presents fully solved problems some quite

often found in major metallurgical operations Physico-chemical Calculations Joseph Knox, 1919 Exam Survival Guide: Physical Chemistry Jochen Vogt, 2017-02-24 A text and exercise book for physical chemistry students This book deals with the fundamental aspects of physical chemistry taught at the undergraduate level in chemistry and the engineering sciences in a compact and practice oriented form Numerous problems and detailed solutions offer the possibility of an in depth reflection of topics like chemical thermodynamics and kinetics atomic structure and spectroscopy Every chapter starts with a recapitulation of important background information before leading over to representative exercises and problems Detailed descriptions systematically present and explain the solutions to the problems so that readers can carefully check their own solutions and get clear cut introductions on how to approach similar problems systematically The book addresses students at the upper undergraduate level as well as tutors and teachers It is a rich source of exercises for exam preparation and can be used alongside classical textbooks Furthermore it can serve teachers and tutors for the conception of their lessons Its well thought through presentation structure and design make the book appeal to everybody who wants to succeed with the physical chemistry lessons and exercises **Numerical Chemistry** , **Chemical Thermodynamics** M L

McGlashan, 2007-10-31 Specialist Periodical Reports provide systematic and detailed review coverage of progress in the major areas of chemical research Written by experts in their specialist fields the series creates a unique service for the active research chemist supplying regular critical in depth accounts of progress in particular areas of chemistry For over 80 years the Royal Society of Chemistry and its predecessor the Chemical Society have been publishing reports charting developments in chemistry which originally took the form of Annual Reports However by 1967 the whole spectrum of chemistry could no longer be contained within one volume and the series Specialist Periodical Reports was born The Annual Reports themselves still existed but were divided into two and subsequently three volumes covering Inorganic Organic and Physical Chemistry For more general coverage of the highlights in chemistry they remain a must Since that time the SPR series has altered

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**Student Solutions Manual** Darrell Ebbing,2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions **Specific Intermolecular Interactions of Organic Compounds** Alexei K. Baev,2012-01-13 Prof Baev presents in his book the development of the thermodynamic theory of specific intermolecular interactions for a wide spectrum of organic compounds ethers ketones alcohols carboxylic acids and hydrocarbons The fundamentals of an unconventional approach to the theory of H bonding and specific interactions are formulated based on a concept of pentacoordinate carbon atoms New types of hydrogen bonds and specific interactions are substantiated and on the basis of the developed methodology their energies are determined The system of interconnected quantitative characteristics of the stability of specific intermolecular interactions is presented The laws of their transformations are discussed and summarized The new concept of the extra stabilizing effect of isomeric methyl groups on the structure and stability of organic molecules is introduced and the destabilization action on specific interactions is outlined Sessional Papers Great Britain. Parliament. House of Commons,1902 Final Report ... Great Britain. Royal College of Science Departmental Committee,1906 **Proceedings** American Society for Engineering Education,Society for the Promotion of Engineering Education (U.S.),1925 **Proceedings of the ... Annual Meeting** Society for the Promotion

of Engineering Education (U.S.). Annual Meeting,1925      *The Journal of Engineering Education* ,1924      **Engineering**  
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