

## Thermochemistry Review - Answers

1. Using the data below, calculate the amount of energy required to melt a 16.2 g aluminum pop can. Assume that the aluminum is at an initial temperature of 20°C.

- Molar heat of fusion of Al(s) = 10.7 kJ/mol
- Melting point of Al(s) = 660°C
- Heat capacity of Al(s) = 0.900 J/g°C

$$\text{mol} = 16.2 \text{ g} \div 26.98154 \text{ g/mol} = 0.600 \text{ mol}$$

$$\Delta H = (20^\circ\text{C} \rightarrow 660^\circ\text{C}) + (\text{melt})$$

$$\Delta H = (mc\Delta T) + (n\Delta H_{\text{fus}})$$

$$\Delta H = (16.2 \text{ g})(0.900 \text{ J/g}^\circ\text{C})(640^\circ\text{C}) + (0.600 \text{ mol})(10.7 \text{ kJ/mol})$$

$$\Delta H = 9331.2 \text{ J} + 6.42 \text{ kJ}$$

$$\Delta H = 9.33 \text{ kJ} + 6.42 \text{ kJ}$$

$$\Delta H = 15.8 \text{ kJ}$$

2. Given the thermochemical equations:



Calculate the  $\Delta H$  for the reaction:  $\text{CH}_3\text{OH}(l) + \frac{3}{2}\text{O}_2(g) \rightarrow \text{CO}_2(g) + 2\text{H}_2\text{O}(l)$



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## **Table of Contents Thermochemistry Review And Answers**

1. Understanding the eBook Thermochemistry Review And Answers
  - The Rise of Digital Reading Thermochemistry Review And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Review And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Thermochemistry Review And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Review And Answers
  - Personalized Recommendations
  - Thermochemistry Review And Answers User Reviews and Ratings
  - Thermochemistry Review And Answers and Bestseller Lists

5. Accessing Thermochemistry Review And Answers Free and Paid eBooks
  - Thermochemistry Review And Answers Public Domain eBooks
  - Thermochemistry Review And Answers eBook Subscription Services
  - Thermochemistry Review And Answers Budget-Friendly Options
6. Navigating Thermochemistry Review And Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Thermochemistry Review And Answers Compatibility with Devices
  - Thermochemistry Review And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Thermochemistry Review And Answers
  - Highlighting and Note-Taking Thermochemistry Review And Answers
  - Interactive Elements Thermochemistry Review And Answers
8. Staying Engaged with Thermochemistry Review And Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Thermochemistry Review And Answers
9. Balancing eBooks and Physical Books Thermochemistry Review And Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Thermochemistry Review And Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Thermochemistry Review And Answers
  - Setting Reading Goals Thermochemistry Review And Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermochemistry Review And Answers
  - Fact-Checking eBook Content of Thermochemistry Review And Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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