



A. Thermochemistry

Theory

1. Explain the difference between kinetic and potential energy using an analogy.
2. Explain the difference between heat, temperature, and thermal energy using an analogy.
3. What are the postulates of the kinetic molecular theory?
4. Describe the difference between an open, closed, and isolated system. Include what can and cannot be exchanged in each case.
5. How do you determine whether a reaction is endothermic or exothermic? Explain two ways.
6. What is calorimeter?
7. Describe the parts of a basic coffee-cup calorimeter, and compare it to a bomb calorimeter.
8. How do the two laws of thermodynamics relate to this unit?
9. How does the enthalpy of physical changes compare to chemical changes and nuclear reactions?
10. What is the difference between nuclear fission and nuclear fusion?
11. Examine a heating curve for water – explain what is occurring at each point along the graph.
12. What assumptions are often made in calorimetry calculations?
13. How are q and ΔH related?
14. What is the difference between ΔH and ΔH_{rxn} ?
15. What side of the thermochemical reaction does the ΔH value appear in an endothermic reaction? Exothermic?
16. Compare potential energy diagrams for endothermic and exothermic reactions.
17. How do energy changes compare when bonds break and when bonds are formed?
18. Why do triple bonds have higher bond energies than single bonds?
19. State Hess's Law.
20. When is a substance in its "standard state"?

Problems

1. In a coffee-cup calorimeter, a student mixes 1.60 g ammonium nitrate with 75.0 g water at an initial temperature of 25.00°C. After dissolution of the salt, the final temperature of the calorimeter contents is 23.34°C. Assuming the solution has the same heat capacity and density as liquid water, calculate the enthalpy change for the dissolution of ammonium nitrate. (+26.0 kJ/mol)
2. In a calorimeter, 60.0 mL of 0.700 mol/L sodium hydroxide solution was neutralized with 40.0 mL of excess sulfuric acid solution. The temperature increased by 5.6°C. Calculate the molar enthalpy of neutralization for sodium hydroxide. (-46.0 kJ/mol)
3. Use bond energies to predict the enthalpy of reaction for $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{CO}(\text{g}) + 3 \text{H}_2(\text{g})$. Write the thermochemical equation and draw a potential energy diagram for this reaction. (218 kJ)
4. Use bond energies to predict the enthalpy of reaction for $\text{H}_2\text{O}_2(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$. Write the thermochemical equation and draw a potential energy diagram for this reaction. (-101.5 kJ)
5. Use Hess's law to determine the enthalpy of reaction for $\text{SO}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{SO}_3(\text{g})$ using the following equations:
 $\text{S}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{SO}_2(\text{g}) \quad \Delta H = -1676 \text{ kJ} \quad (-100.9 \text{ kJ})$
 $2 \text{S}(\text{s}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{SO}_3(\text{g}) \quad \Delta H = -824 \text{ kJ}$
6. Use Hess's Law and the standard enthalpies of formation to determine the enthalpy of reaction for:
 $2 \text{NH}_3(\text{g}) + 3 \text{O}_2(\text{g}) + \text{CH}_4(\text{g}) \rightarrow 2 \text{HCN}(\text{g}) + 6 \text{H}_2\text{O}(\text{g}) \quad (-1014.4 \text{ kJ})$

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