

38. At equilibrium which is correct :- [AIIMS-2012]
 (1) $\Delta G = 0$ (2) $\Delta S = 0$ (3) $\Delta H = 0$ (4) $\Delta G^\circ = 0$
39. Bond dissociation energy of CH_4 is 360 kJ/mol and C_2H_6 has 620 kJ/mol. Then bond dissociation energy of C-C bond is :- [AIIMS-2012]
 (1) 170 kJ/mol (2) 50 kJ/mol (3) 80 kJ/mol (4) 220 kJ/mol
40. Which thermodynamic parameter is not a state function :- [AIIMS-2013]
 (1) q at constant pressure (2) q at constant volume
 (3) w at adiabatic (4) w at isothermal
41. $\Delta S_{\text{total}} = -40 \text{ kJ/mol} \times \text{K}$
 $\Delta H_{\text{sys}} = 2000 \text{ kJ/mol}$
 $T = 400 \text{ K}$
 Find out value of ΔS_{system} ? [AIIMS-2014]
 (1) $-35 \text{ kJ/mol} \times \text{K}$ (2) $-5 \text{ kJ/mol} \times \text{K}$
 (3) $-40 \text{ kJ/mol} \times \text{K}$ (4) $-5 \text{ kJ/mol} \times \text{K}$
42. Heat of sublimation of Na = x J [AIIMS-2014]
 Ionization energy of Na = y J
 Bond dissociation energy of $\text{Cl}_2 = a \text{ J}$
 e^- gain enthalpy of Cl = - b J
 Enthalpy of formation of NaCl = - c J
 Then what will be lattice energy of NaCl_{cr}
 (1) $b + \left(x + y + c - \frac{a}{2}\right)$ (2) $b - \left(x + y + c - \frac{a}{2}\right)$
 (3) $b - \left(\frac{a}{2} + x + y + c\right)$ (4) $c - \frac{a}{2} + x + y + c$
43. Using the Gibbs energy change, $\Delta G^\circ = +63.3 \text{ kJ}$, for the following reaction,
 $\text{Ag}_2\text{CO}_3 \rightleftharpoons 2\text{Ag}^+ (\text{aq}) + \text{CO}_3^{2-} (\text{aq})$
 the K_{sp} of $\text{Ag}_2\text{CO}_3(\text{s})$ in water at 25°C is :- [AIPMT-2014]
 ($R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$)
 (1) 3.2×10^{-26} (2) 8.0×10^{-12} (3) 2.9×10^{-3} (4) 7.9×10^{-2}
44. For the reaction : $\text{N}_2\text{O}_4(\text{l}) \longrightarrow 2\text{NO}_2(\text{g})$
 $\Delta U = 2.1 \text{ k cal}$, $\Delta S = 20 \text{ cal K}^{-1}$ at 300 K
 Hence ΔG is :- [AIPMT-2014]
 (1) 2.7 k cal (2) -2.7 k cal (3) 9.3 k cal (4) -9.3 k cal

Thermochemistry And Thermodynamics Questions And

K Morrison



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General Chemistry S. Ekambaram, 2012 General Chemistry presents the fundamental concepts of general chemistry in a precise and comprehensive manner for undergraduate students of chemistry and life science at all Indian universities Adhering strictly to the UGC curriculum the contents are written in a simple and lucid language enriched with a large number of examples and illustrations **101 Ways to Score Higher on Your MCAT** Marti Anne Maguire, Paula Stiles, 2010 Contains over one hundred tips and suggestions to improve SAT scores with step by step instructions to write better essays and reviews of grammar usage and math concepts details a test taking strategy that encourages time management and educated guessing and provides resources for practice tests *Chemistry for Degree Students B.Sc. (Honours) Semester II, 1/e (As per CBCS)* Madan R.L., 2022 This textbook has been designed to meet the needs of B Sc Honours Second Semester students of Chemistry as per the UGC Choice Based Credit System CBCS Maintaining the traditional approach to the subject this textbook lucidly explains the basics of Organic and Physical Chemistry Important topics such as alkanes alkenes alkynes stereochemistry aliphatic hydrocarbons thermochemistry chemical thermodynamics and chemical equilibrium are aptly discussed to give an overview of organic and physical chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures *Chemical Thermodynamics* M L McGlashan, 2007-10-31 Specialist Periodical Reports provide systematic and detailed review coverage of progress in the major areas of chemical research Written by experts in their specialist fields the series creates a unique service for the active research chemist supplying regular critical in depth accounts of progress in particular areas of chemistry For over 80 years the Royal Society of Chemistry and its predecessor the Chemical Society have been publishing reports charting developments in chemistry which originally took the form of Annual Reports However by 1967 the whole spectrum of chemistry could no longer be contained within one volume and the series Specialist Periodical Reports was born The Annual Reports themselves still existed but were divided into two and subsequently three volumes covering Inorganic Organic and Physical Chemistry For more general coverage of the highlights in chemistry they remain a must Since that time the SPR series has altered according to the fluctuating degree of activity in various fields of chemistry Some titles have remained unchanged while others have altered their emphasis along with their titles some have been combined under a new name whereas others have

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