

THERMOCHEMISTRY CALCULATIONS WORKSHEET 1

Using reaction equation ratios

- How much heat will be released when 6.44 g of Sulfur reacts with excess O_2 according to the following equation?
 $2S + 3O_2 \rightarrow 2SO_2 \quad \Delta H^\circ = -791.4 \text{ kJ}$
- How much heat will be released when 4.72 g of Carbon reacts with excess O_2 according to the following equation?
 $C + O_2 \rightarrow CO_2 \quad \Delta H^\circ = -393.5 \text{ kJ}$
- How much heat will be absorbed when 38.2 g of Bromine reacts with excess H_2 according to the following equation?
 $H_2 + Br_2 \rightarrow 2HBr \quad \Delta H^\circ = 72.80 \text{ kJ}$
- How much heat will be released when 1.48 g of Chlorine reacts with excess phosphorus according to the following equation?
 $2P + 3Cl_2 \rightarrow 2PCl_3 \quad \Delta H^\circ = -886 \text{ kJ}$
- How much heat will be released when 4.77 g of ethanol (C_2H_5OH) reacts with excess O_2 according to the following equation?
 $C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O \quad \Delta H^\circ = -1366.7 \text{ kJ}$
- How much heat will be absorbed when 13.7 g of Nitrogen reacts with excess O_2 according to the following equation?
 $N_2 + O_2 \rightarrow 2NO \quad \Delta H^\circ = -180 \text{ kJ}$
- How much heat will be released when 11.8 g of Iron reacts with excess O_2 according to the following equation?
 $3Fe + 2O_2 \rightarrow Fe_2O_3 \quad \Delta H^\circ = -1120.5 \text{ kJ}$
- How much heat will be released when 18.6 g of Hydrogen reacts with excess O_2 according to the following equation?
 $2H_2 + O_2 \rightarrow H_2O \quad \Delta H^\circ = -571.6 \text{ kJ}$
- How much heat will be transferred when 14.9 g of ammonia reacts with excess O_2 according to the following equation?
 $4NH_3 + 5O_2 \rightarrow 4NO + 6H_2O \quad \Delta H^\circ = -1170 \text{ kJ}$
- How much heat will be transferred when 5.81 g of graphite reacts with excess H_2 according to the following equation?
 $6C(\text{graphite}) + 3H_2 \rightarrow C_6H_6 \quad \Delta H^\circ = 49.03 \text{ kJ}$

Using $\Delta H = mC\Delta T$

- How many kilojoules of heat energy are required to heat all the aluminum (C_p of Al = $.902 \text{ J/g}\cdot^\circ\text{C}$) in a roll of aluminum foil (500.0 g) from room temperature (25.0°C) to the temperature of a hot oven (250.0°C)?
- One way to cool down your cup of coffee is to plunge an ice-cold piece of aluminum into it. Suppose you store a 20.0 g piece of aluminum (C_p of Al = $.902 \text{ J/g}\cdot^\circ\text{C}$) in the refrigerator at 4.40°C and then drop it into your coffee. The coffee temperature drops from 90.0°C to 55.0°C . How many joules of heat energy did the aluminum block absorb? (Ignore the cooling of the cup)
- Suppose you pick up a 16.0-kilogram ball of iron (such as a "shot-put" ball at a track event). The iron (C_p = $.451 \text{ J/g}\cdot^\circ\text{C}$) ball has the same temperature as the atmosphere on a cool day say 16.0°C . How many kilojoules of heat energy must the iron ball absorb to reach the temperature of your body (37.0°C)?
- The specific heat of silver is $0.24 \text{ J/g}\cdot^\circ\text{C}$. How much heat in joules must be added to a silver block of mass 86.0 g to raise its temperature by 9.00°C ?
- An 18.7 g sample of platinum metal increases in temperature by 2.30°C when 5.70 J of heat are added. What is the specific heat of platinum?

Answers: 1) -79.4kJ 2) -155kJ 3) 17.4kJ 4) -3.75kJ 5) -142kJ 6) -88.1kJ 7) -78.7kJ 8) -2660kJ 9) -256kJ 10) 3.96kJ
11) 101kJ 12) 912J 13) 152kJ 14) 186J 15) $.132 \text{ J/g}\cdot^\circ\text{C}$

Thermochemistry Calculations Answers Keys

James G. Anderson



Thermochemistry Calculations Answers Keys:

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Pearson Edexcel A Level Chemistry (Year 1 and Year 2) Andrew Hunt, Graham Curtis, Graham Hill, 2019-07-15 Develop and assess your students knowledge and skills throughout A level with worked examples practical assessment guidance and differentiated end of topic questions in this updated all in one textbook for Years 1 and 2 Combining everything your students need to know for the Pearson Edexcel A level Chemistry specification this revised textbook will Identify the level of your students understanding with diagnostic questions and a summary of prior knowledge at the start of the Student Book Provide support for all 16 required practicals with various activities and questions along with a Practical chapter covering procedural understanding and key ideas related to measurement Improve mathematical skills with plenty of worked examples including notes on methods to help explain the strategies for solving each type of problem Offer plenty of practice with Test yourself questions to help students assess their understanding and measure progress Encourage further reading and study with short passages of extension material Develop understanding with free online access to Test yourself answers and an extended glossary

Understanding General Chemistry Atef Korchef, 2022-03-07 Understanding General Chemistry details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter Written in an easy to understand format with helpful pedagogy to fuel learning the book features main objectives at the beginning of each chapter get smart sections and check your reading section at the end of each chapter The text is filled with examples and practices that illustrate the concepts at hand In addition a summary and extensive MCQs exercises and problems with the corresponding answers and explanations are readily available Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes Offers answers and comments alongside sample problems enabling students to self evaluate their skill level

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STOICHIOMETRY AND PROCESS CALCULATIONS K.V. NARAYANAN ,B. LAKSHMIKUTTY,2016-12-01 Designed as a textbook for the undergraduate students of chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering and safety engineering the chief objective of the book is to prepare students to make analysis of chemical processes through calculations and to develop systematic problem solving skills in them The text presents the fundamentals of chemical engineering operations and processes in a simple style that helps the students to gain a thorough understanding of chemical process calculations The book deals with the principles of stoichiometry to formulate and solve material and energy balance problems in processes with and without chemical reactions With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations The book is supplemented with Solutions Manual for instructors containing detailed solutions of all chapter end unsolved problems NEW TO THE SECOND EDITION Incorporates a new chapter on Bypass Recycle and Purge Operations Comprises updations in some sections and presents new sections on Future Avenues and Opportunities in Chemical Engineering Processes in Biological and Energy Systems Contains several new worked out examples in the chapter on Material Balance with Chemical Reaction Includes GATE questions with answers up to the year 2016 in Objective type questions KEY FEATURES SI units are used throughout the book All basic chemical engineering operations and processes are introduced and different types of problems are illustrated with worked out examples Stoichiometric principles are extended to solve problems related to bioprocessing environmental engineering etc Exercise problems more than 810 are organised according to the difficulty level and all are provided with answers **Class 8-12 Chemistry Questions and Answers PDF** Arshad Iqbal, The Class 8 12 Chemistry Quiz Questions and Answers PDF Grade 8 12 Chemistry Competitive Exam Questions Chapter 1 15 Practice Tests Chemistry Textbook Questions for Beginners includes Questions to solve problems with hundreds of class questions Class 8 12 Chemistry Questions and Answers PDF book covers basic concepts and analytical assessment tests Class 8 12 Chemistry Quiz PDF book helps to practice test questions from exam prep notes The Grade 8 12 Chemistry Quiz Questions and Answers PDF eBook includes Practice material with verbal quantitative and analytical past papers questions Class 8 12 Chemistry Questions and Answers PDF Free download chapter 1 a book to review textbook questions on chapters Molecular structure acids and bases atomic structure bonding chemical equations descriptive

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Mikulecky,Katherine Brutlag,Michelle Rose Gilman,Brian Peterson,2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you ll build your chemistry skills and improve your performance both in and out of the science lab You ll see how to

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AQA AS Chemistry: Unit 1 Foundation Chemistry Student Unit Guide Ebook Margaret Cross,2008-09-26 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements summarises the relevant unit content and includes a series of specimen questions and answers There are three sections to each guide

Introduction includes advice on how to use the guide an explanation of the skills being tested by the assessment objectives an outline of the unit or module and depending on the unit suggestions for how to revise effectively and prepare for the examination questions Content Guidance provides an examiner s overview of the module s key terms and concepts and identifies opportunities to exhibit the skills required by the unit It is designed to help students to structure their revision and make them aware of the concepts they need to understand the exam and how they might analyse and evaluate topics Question and Answers sample questions and with graded answers which have been carefully written to reflect the style of the unit All responses are accompanied by commentaries which highlight their respective strengths and weaknesses giving students an insight into the mind of the examiner

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