

Answers - Exercises (Thermochemistry)

Q1: Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15g of copper from 25°C to 60°C. (Ans : $c = 0.39 \text{ J/g.K}$)

Answer:

$$q = mc\Delta T$$

$$q = 204.75 \text{ J}$$

$$m = 15\text{g}$$

$$\Delta T = (60 - 25) = 35^\circ\text{C}$$

$$c = \frac{q}{m\Delta T}$$

$$= \frac{204.75 \text{ J}}{(15\text{g})(35^\circ\text{C})}$$

$$= 0.39 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$$

Q2: A chemist burns 0.8650 g of graphite (a form of carbon) in a new bomb calorimeter, and CO_2 forms. If 393.5 kJ of heat is released per mole of graphite and T increases 2.613 K, what is the capacity of the bomb calorimeter? (Ans: 10.84 kJ/K)

Answer:



$$\text{Mass of Carbon (Graphite)} = 0.8650 \text{ g}$$

$$\text{Atomic mass of Carbon} = 12 \text{ g/mol}$$

$$\text{Mol of graphite} = \text{mass/atomic mass} = 0.8650 \text{ g/} 12\text{g mol}^{-1} = 0.072 \text{ mol}$$

$$\text{Heat of combustion, } \Delta H_{\text{comb}} = -393.5 \text{ kJ/mol}$$

$$\text{Temperature change, } \Delta T = 2.613 \text{ K}$$

$$\text{Calorimeter heat capacity, } C_{\text{cal}} = ?$$

$$\Delta H_{\text{comb}} = \frac{q_{\text{rxn}}}{n} = \frac{C_{\text{cal}} \times \Delta T}{n}$$

$$393.5 = \frac{C_{\text{cal}}(2.613)}{0.072}$$

$$C_{\text{cal}} = 10.84 \text{ kJ/K}$$

Thermochemistry Exercises Answers

Maka Aleksishvili, Shota Sidamonidze



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Solutions Manual to Accompany Elements of Physical Chemistry C. A. Trapp, M. P. Cady, 2007 The Solutions manual to accompany Elements of Physical Chemistry 4e contains full worked solutions to all end of chapter exercises featured in the book

Ions in Solution J Burgess, 1999-10-01 This outline of the principles and chemical interactions in inorganic solution chemistry delivers a course module in an area of considerable complexity Problems with solutions and tutorial hints to test comprehension have been added as a feature to check readers understanding and assist self study Exercises and projects are also provided to help readers deepen and extend their knowledge and understanding Inorganic solution chemistry is treated thoroughly Emphasis is placed upon NMR UV VIS IR Raman spectroscopy X ray diffraction and such topics as acid base behaviour stability constants and kinetics

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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each

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Problem Exercises for General Chemistry G.Gilbert Long,Forrest C. Hentz,1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT Problems in Chemical Thermodynamics with Solutions Maka Aleksishvili,Shota Sidamonidze,2002 The methods of chemical thermodynamics are effectively used in many fields of science and technology Mastering these methods and their use in practice requires profound comprehension of the theoretical questions and acquisition of certain calculating skills This book is useful to undergraduate and graduate students in chemistry as well as chemical thermal and refrigerating technology it will also benefit specialists in all other fields who are interested in using these powerful methods in their practical activities

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Langley, 2022-06-12 A unique and effective way to learn Chemistry updated with the latest instruction and review Must Know High School Chemistry provides a fresh approach to learning As part of our Must Know series this new edition makes sure what you really need to know is clear up front Rather than starting with goals to be met chapters begin by telling you the most important concepts about the topic at hand and then show you exactly how these concepts help you accomplish your goals Written by expert chemistry educators Must Know High School Chemistry Second Edition provides updated lesson content and useful examples to help clarify each topic Every chapter closes with reinforcing exercises to get you the practice you need to gain confidence New features to this edition focus on extra support and helping you avoid common mistakes In the end you get everything you need to build your chemistry skills quickly and painlessly Features More than 350 practice questions that parallel what you will find in your classwork and on exams Bonus app that includes 100 flashcards to reinforce concepts Extra Help and Easy Mistake features put the emphasis on how to improve and what pitfalls to avoid Chemistry topics aligned to national and state educational standards Practical examples throughout and an answer key with explanations make sure you understand the topics Conversational writing style and informative IRL In Real Life and BTW By the Way sidebars A special section for teachers with tips and strategies on teaching the material and content specific links and resources **Chemistry-vol-I** Dr Vijay Sarda, Dr A C Handa, Dr K K Arora, A text book on Chemistry Chemistry:

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