CHEMISTRY 161 THERMO WORKSHEET FALL 2015

- A system absorbs 52.5 J of energy as it expands to a volume of 50.0 L. If the external pressure is 0.800 atm and the change in internal energy is -112 j, calculate the initial volume.

- How much glucose must be metabolized to completely evaporate 1.00 g of water? The heat of combustion of glucose is 2803 kj/mole.
- The standard enthalpy of formation of NH₃ is -46.1 kJ/mole. What is ΔH°_{cov} for the following reactions?
 a. N₂(g) + 3 H₂(g) → 2NH₃(g)
 b. NH₃(g) → N₂(g) + 3 H₂(g)
- The heat of combustion of pentane (C₅H₁₂) is -3535 kj/mole
 a. Find the heat given off when 1.00 kj of pentane is burned.
 - b. How many grams of pentane must be burned to increase the temperature of 1.00 kg of water from 20.0°C to 90.0°C?
- 6. Brass has a density of 8.40 g/cm³ and a specific heat of 0.385 j/°C. A 15.2 cm³ piece of brass at an initial temperature of 163°C is dropped into an insulated container with 150.0 g H₂O at 22.4°C. What will be the final temperature of the brass-water mixture?

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